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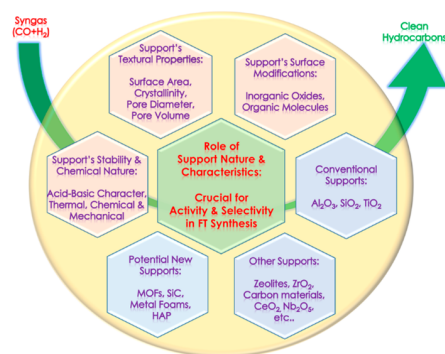
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Effect of the Support and Its Surface Modifications in Cobalt-Based Fischer–Tropsch Synthesis

Rajesh Munirathinam,*¹ Doan Pham Minh, and Ange Nzihou

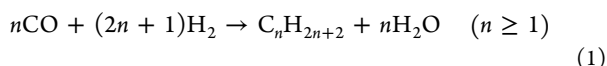
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ABSTRACT: Support effects in heterogeneous catalysis are evolving as an important field of investigation to optimize catalyst properties. The cobalt-based Fischer–Tropsch (FT) catalysts usually consist of metallic cobalt nanocrystallites dispersed on a support material. The present review surveys the progress that has been made over the last couple of decades in the area of the effect of the support and its surface modifications in cobalt-based FT synthesis. Different catalyst supports such as alumina, silica, titania, niobia, zirconia, zeolite, ceria, carbon-based materials, silicon carbide, aluminum phosphate, hydrotalcite, metal–organic frameworks, and metal foams are discussed and compared with classical supports like alumina or silica wherever permitted. Properties such as metal–support interactions, the support-induced size and morphology of the cobalt nanocrystallites (textural properties of the support), changes in the electronic properties of the cobalt clusters, and the acid/base nature of the support are examined, and wherever possible the activity and/or selectivity in FT synthesis is discussed. This review also summarizes findings on new and promising supports for FT catalysts. Guidance for support modifications and choice of the support as functions of the product selectivity in the FT process is also proposed.



1. INTRODUCTION

Limited petroleum reserves have called for the strengthening of research activities on Fischer–Tropsch (FT) synthesis.¹ It aims to produce synthetic fuels and value-added chemicals from synthesis gas (syngas, containing mainly CO and H₂) obtained from non-petroleum resources such as coal, natural gas, biomass, and wastes. Cobalt² and iron³ are the most widely used metal catalysts in FT synthesis. Cobalt catalysts are preferred over iron catalysts for the low-temperature FT synthesis because of their increased stability, significant resistance toward oxidation by the steam formed during the course of the FT reaction, poor activity in the water-gas shift reaction, and increased selectivity for long-chain linear paraffins.^{4,5} The overall FT reaction that yields straight-chain alkane (paraffin) products is represented by the simplified eq 1:



Industrially used cobalt-based FT catalysts (Co catalysts) with cobalt loadings in the range of 10–30 wt % are usually supported on alumina (Al₂O₃), silica (SiO₂), or titania (TiO₂).^{6,7} The use of other supports such as carbon-based supports,⁸ mesoporous materials (SBA-15, MCM-41),⁹ zirconia (ZrO₂),¹⁰ niobia (Nb₂O₅),¹¹ ceria (CeO₂),¹² zeolites,¹³ and SiC¹⁴ has also been reported in the literature. The Co catalysts are often promoted with noble metals (Ru, Pt, Re, and Pd in the range of 0.05–1 wt %) to enhance the reducibility of cobalt oxide particles, to improve cobalt dispersion, inhibit catalyst deactivation, increase the number of hydrogen activation sites, and consequently augment the

intrinsic reactivity of active sites.¹⁵ In addition, transition metal oxides such as zirconia, chromia, and manganese oxides and some rare-earth metal oxides such as lanthanum oxide, ceria, and thorium oxide have also been employed as structural promoters (in the range of 1–10 wt %) to improve the dispersion of the active phase and its stability, consequently leading to increased catalytic activity.¹⁶ Because of the overlapping effects of structural promoters that influence the formation and stability of the active phase, textural promoters, and electronic promoters that affect elementary steps involved in each turnover on the catalyst, it is difficult to precisely define the function of a promoter.¹⁶

The support is a porous refractory oxide that is thermally stabilized at elevated temperatures to secure the strength and stability of the material. Generally, increasing the calcination temperature of the support leads to the formation of a more inert surface with reduced surface area, creation of larger pores, and condensation of surface hydroxyl groups. The effect of the support is an important field of investigation in heterogeneous catalysis.¹⁷ The primary functions of the support are to disperse the active metal, yield a high surface area of the catalytic phase, and provide mechanical strength and thermal stability to the active phase. Furthermore, the support's chemical nature also influences several important properties such as the reducibility of metal precursors, heat and mass transfer in an exothermic or diffusion-limited reaction, the

electronic state of the active metal, diffusion of reagents and reaction products, and the stability of the active phase against morphological change during the reaction.^{18,19}

The textural properties of the support, such as the surface area, pore structure, and surface chemistry, greatly affect cobalt dispersion. For example, high-surface-area supports provide high cobalt dispersion and most likely give rise to highly active FT catalysts, with the condition that there are no detrimental metal–support interactions that yield catalytically inactive species.²⁰ The acidity of the support is known to play a role in determining the selectivity for methane, short-chain hydrocarbons, or long-chain hydrocarbons.²¹ Since the FT reaction is exothermic, the support also plays an important role in dissipating heat and thereby reducing the temperature gradients, especially in fixed-bed reactors. Therefore, the choice of the carrier material is of vital importance in FT synthesis.²² Comparison of the site activities of Co catalysts on different supports in FT synthesis is not straightforward, most likely because of the great variability of intrinsic textural properties among the different support materials and catalyst testing conditions. For example, a study of CO hydrogenation over cobalt catalysts supported on different carriers (3 wt % Co loading) at 1 atm and 225 °C displayed specific activities for CO hydrogenation in the following order: Co/TiO₂ > Co/SiO₂ > Co/Al₂O₃ > Co/C > Co/MgO.²² On the contrary, another study² showed that the turnover frequency for CO conversion on Co catalysts (dispersion range, i.e., fraction of surface atoms, of 0.01–0.12) was independent of the support (Al₂O₃, SiO₂, TiO₂, SiO₂-modified TiO₂, and MgCr₂O₄) under typical FT reaction conditions (20 bar and 200 °C). In a different study,²³ the cobalt site activity for CO conversion and methane selectivity remained unaffected by the identity of the support (Al₂O₃, SiO₂, and TiO₂); however, modifying the support with Y₂O₃, MgO, or ZnO (by coating) reduced the site activity, thus demonstrating the importance of the nature of the interface between the support and the metal.

Research concerning the development of new catalyst carriers and their surface modifications to support cobalt has been active over the last couple of decades to decode the effect of the support in FT synthesis. Although results about support effects have been summarized partly in general reviews,^{5,19,24,25} a general review specifically on this topic is still missing, probably because of the complexity of the effect of the support on the catalytic behavior and the difficulty of comparing supports subjected to different experimental conditions. Rytter and Holmen²⁶ attempted to discuss the impact of alumina and aluminates in FT synthesis, but the discussion was limited only to γ -Al₂O₃, α -Al₂O₃, magnesium aluminate, and nickel aluminate.

During the last decades, much research work has been devoted to the valorization of biomass and wastes using thermal conversion processes (pyrolysis, gasification). Biomass and waste-derived syngas generally contain complex and variable pollutants, including particulate matter, tars, sulfur-containing compounds, nitrogen-containing compounds, halide-containing compounds, and even metals. Purification of this syngas is difficult because of the variable nature of biomass and wastes as pyrolysis/gasification feedstocks. Thus, the valorization of this biomass- and waste-derived syngas via catalytic processes such as FT synthesis needs new catalysts that are highly resistant to pollutants.^{27,28} A survey of the latest developments on FT synthesis catalysts appears pertinent at this point.

In this review, we attempt to examine the state of the art on support effects, including most of the supports used in FT synthesis from research publications over the last two decades. Supports that have been surface-modified by the introduction of inorganic oxide and organic precursors before impregnation of the cobalt metal precursors are also surveyed in this review.

2. ALUMINA-SUPPORTED CATALYSTS

Alumina (Al₂O₃) is the most commonly used carrier material in industry because of its low cost, good thermal and mechanical stability, and ability to yield a highly dispersed active phase. Al₂O₃ can be obtained in different crystallographic forms, i.e., γ -Al₂O₃, θ -Al₂O₃, and α -Al₂O₃ derived from boehmite or η -Al₂O₃ resulting from bayerite.²⁹ γ -Al₂O₃ is the most employed support in FT catalysts. Al₂O₃ is known to interact with cobalt ions and yield a well-dispersed active phase. As a result of metal–support interactions that arise from the diffusion of cobalt ions into alumina lattice sites of octahedral or tetrahedral geometry (forming CoAl₂O₄ species) during calcination, the reducibility of the catalysts is hindered, and a portion of the metal ions remain catalytically inactive.³⁰ Complete diffusion of cobalt ions into lattice sites (tetrahedral) is attained only at low concentrations (i.e., ca. ≤ 2 wt % Co), and the percentage of cobalt ions interacting with the carrier decreases with increasing cobalt loading and decreasing calcination temperature.³¹ A study of cobalt catalysts supported on Al₂O₃, SiO₂, and TiO₂ showed that the metal–support interactions affect the reduction of cobalt species, and the strengths of such interactions for different supports decrease in the order Al₂O₃ > TiO₂ > SiO₂.³² The varied metal–support interactions give rise to differences in reducibility and dispersion of the active phase on different supports, consequently affecting the specific activity of cobalt-based FT catalysts. Cobalt reducibility is one of the major limitations of Al₂O₃-supported cobalt FT catalysts that arise from the detrimental metal–support interactions. It is circumvented in most cases by promotion with a noble metal such as Ru, Re, Pt, or Pd, which has been discussed in several reviews.^{5,15,16} In the following sections, the effects of the textural properties and crystalline phase of alumina, surface modification of alumina by different metal oxides and organic molecules, and the acid/base properties of alumina on the catalytic activity/selectivity in FT synthesis are discussed.

2.1. Effect of the Textural Properties and Crystalline Phase of the Alumina Support. Textural properties such as surface area, pore volume, pore size distribution, and crystalline phase of alumina are known to influence the activity/selectivity in Co-based FT synthesis through varied degrees of Co reducibility or formation of a cobalt aluminate phase.^{24,25} The support pore size plays a role in controlling the size of cobalt (Co₃O₄) particles, with the formation of small particles in narrow pores and large particles in wide pores.^{33–36} Borg et al.³⁵ showed that the degree of Co reduction increased with increasing catalyst pore size and Co₃O₄ particle size. Furthermore, they observed positive correlations between the cobalt particle size and the C₅₊ selectivity and between the catalyst pore size and the C₅₊ selectivity, with small variations in the cobalt site-time yields related to the chemical purity of the catalyst samples. Low-surface-area alumina (α -Al₂O₃) was found to display better C₅₊ selectivity compared with high-surface-area alumina (γ -Al₂O₃).^{36,37} A macroporous-structured (MPS) Al₂O₃ support synthesized from Al₂O₃ nanoparticles using polystyrene beads as a sacrificial agent seemed to

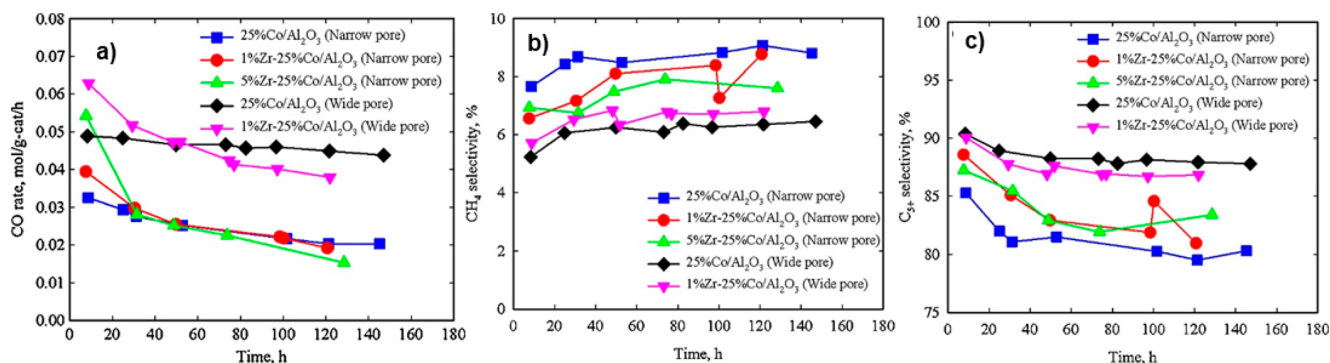


Figure 1. Effect of Zr loading and pore size on (a) CO rate, (b) CH₄ selectivity, and (c) C₅₊ selectivity. Reaction conditions: 220 °C, 22 bar, H₂/CO = 2.1, X_{CO} = 49–51%. Adapted with permission from ref 49. Copyright 2014 Elsevier.

combine the advantages of α -Al₂O₃ and γ -Al₂O₃ as support materials, with activity and selectivity comparable to those of Co/ γ -Al₂O₃ and Co/ α -Al₂O₃, respectively.³⁸ Vosoughi et al.³⁹ demonstrated the synthesis of mesoporous alumina using Pluronic F127 as a structure-directing agent that contributed to a high surface area and large pore volume and pore diameter of the support. The structure of the thus-prepared support collapsed upon impregnation with aqueous cobalt nitrate solution, but shifting the impregnation medium from aqueous to organic (ethanol or acetone) reinstated the textural stability of the mesoporous alumina. The cobalt catalyst supported on mesoporous alumina revealed higher CO conversion (8.3%), higher hydrocarbon yield (18%), and lower CH₄ selectivity (2.6%) compared with its counterpart supported on γ -Al₂O₃. In a more detailed study, different alumina phases (δ -, θ -, and α -Al₂O₃) prepared by thermal treatment of small-pore (11–19 nm) and medium-pore (21–36 nm) γ -Al₂O₃ were charged with a constant Co loading (12 wt %) containing cobalt particle sizes in the range of 8–10 nm.⁴⁰ It was found that the medium-pore catalysts were reduced at lower temperatures and also exhibited higher selectivity for C₅₊ compared with the small-pore catalysts (12 wt % Co, 210 °C, 20 bar, H₂/CO = 2). In addition, higher C₅₊ selectivity was observed for δ - and α -Al₂O₃ phases compared with θ - and γ -Al₂O₃, with α -Al₂O₃ giving the highest selectivity. The site-time yield for cobalt supported on alumina did not vary systematically with the support variations investigated. These studies clearly established the role of textural properties of the support in controlling the particle size, reducibility, activity, and selectivity of the catalyst in FT synthesis.

2.2. Effect of Surface Modifications on the Alumina Support. Modification of the alumina surface is one of the options to weaken the interaction between the alumina support and the cobalt oxides, thereby limiting or suppressing the direct Co–Al interactions and helping to increase the reducibility of cobalt particles. Precoating of cobalt aluminate on the surface of common alumina (γ -Al₂O₃) exhibited higher cobalt specific activity, higher C₅₊ selectivity, higher C₅₊ space-time yields, and lower methane selectivity than conventional catalysts in the FT synthesis.⁴¹ The enhanced catalytic performance was ascribed to the formation of the more active hexagonal close-packed (hcp) Co phase on the aluminate-modified alumina-supported cobalt catalysts, as opposed to the face-centered cubic (fcc) phase, which was supported on unmodified alumina. In addition, the modified catalysts exhibited lower reduction temperatures and a more intense bridged CO adsorption mode that generally yields long-chain

hydrocarbons. In the following sections, both inorganic oxides and organic molecules that have been used to prevent the direct interaction of cobalt with the support are discussed.

2.2.1. Surface Modifications with Inorganic Oxides.

2.2.1.1. Modification with Magnesia and Barium Oxide.

Studies have shown that modification of γ -Al₂O₃ with magnesia (MgO) increased the reducibility of cobalt oxide, positively impacting the catalytic activity in CO hydrogenation.^{42,43} It was found that the addition of small amounts of magnesia to γ -Al₂O₃ improved the activity of cobalt catalysts in FT synthesis. However, larger amounts of magnesia (>0.8 wt %) restrained the reduction of the catalysts because of the formation of a MgO–CoO phase, consequently decreasing the activity and increasing the methane and CO₂ selectivity.⁴² A Mg (10 wt %)-modified α -Al₂O₃ (MgAl₂O₄– α -Al₂O₃) support prepared by incipient wetness impregnation of γ -Al₂O₃ with Mg(NO₃)₂ and calcination in air at 1140 °C for 10 h displayed better mechanical strength compared with unmodified α -Al₂O₃, making it more suitable to operate in slurry or fluidized-bed reactors.⁴⁴ The reduction of CoO to Co⁰ decreased with increasing Mg loading (≥ 1 wt %), consequently impacting negatively both the catalytic activity and selectivity. The negative effect of Mg was attributed to a chemical/site (electronic) effect similar to that observed for alkali metals (Na, K, and Li), in which case the surface H concentration was decreased and CO adsorption and dissociation were increased.⁴⁵

Cobalt catalysts supported on barium modified γ -Al₂O₃ displayed improved cobalt reducibility when barium is added in smaller amounts (≤ 2 wt %) and thus led to more cobalt active sites on the catalyst surface.⁴⁶ These catalysts tested in a continuously stirred tank reactor presented higher CO conversion (ca. 36% versus 32%) and C₅₊ selectivity (ca. 83% versus 80%) compared to unmodified catalyst in FT synthesis (2 wt % Ba, 15 wt % Co, 10 bar, 235 °C and H₂/CO = 2). On the contrary, catalysts with high BaO loadings showed negative effects on the activity and selectivity for high hydrocarbons owing to low cobalt reducibility. Therefore, choosing the right amount of the modifying material is necessary to gain improvement in the catalytic performance.

2.2.1.2. Modification with Zirconia. A cobalt catalyst supported on a zirconium-modified γ -Al₂O₃ support displayed a decrease in the content of the CoAl₂O₄ spinel phase with increasing zirconium loading (0.5–15 wt %) on the γ -Al₂O₃ support, resulting from minimization of the direct metal–support interactions.⁴⁷ Zr addition improved the CO hydrogenation activity and C₅₊ selectivity of the Co/ γ -Al₂O₃ catalyst

in FT synthesis (15 wt % Co, 230 °C, 10 bar, H₂/CO = 2). The enhanced catalytic performance was attributed to an increase in the number and reducibility of cobalt active sites. On the contrary, Rohr et al.⁴⁸ reported that modification of the Al₂O₃ (boehmite phase) support with Zr (3–10 wt %) improved neither the reducibility nor the dispersion of cobalt, which might be ascribed to the different pretreatment conditions of the support and catalyst preparation. Nevertheless, both the activity and selectivity for higher hydrocarbons increased compared with the unmodified catalysts (5–20 wt % Co, 245 °C, 5 bar, H₂/CO = 9). Ma and co-workers⁴⁹ systematically studied the impact of the Zr loading (1–5 wt %) and pore size (10.8 and 25 nm) of the Al₂O₃ support in cobalt-based FT synthesis at a CO conversion level of ca. 50% (25 wt % Co, 220 °C, 22 bar, H₂/CO = 2.1) (Figure 1a). They found that the cobalt site density of wide-pore catalysts increased by 25–55% compared with narrow pore catalysts. The catalytic test results indicated that Zr-modified alumina supports significantly improved the activity per gram of catalyst (g_{cat}), but the catalyst deactivation rate was increased. Furthermore, the pore size (wide or narrow) and loading of Zr promoter (1–5 wt %) did not change the cobalt intrinsic activity. However, addition of Zr (1–5 wt %) decreased the CH₄ selectivity and increased the C₅₊ selectivity for the narrow-pore 25% Co/Al₂O₃ catalyst, whereas the opposite trend was observed with the wide-pore Co catalyst (Figure 1b,c). The differences in the selectivity trends were attributed to the variances in the pore size and cobalt distribution.

2.2.1.3. Modification with Titania. Surface modification of γ -Al₂O₃ with small quantities of titanium (0.2–0.5 wt % Ti–Al₂O₃) was found to reduce direct detrimental metal–support interactions and stabilize the grain boundary of Al₂O₃–TiO₂, consequently altering the density of surface acidic sites, which in turn led to a homogeneous size distribution of cobalt clusters with increased reducibility.⁵⁰ Cobalt catalysts supported on Ti–Al₂O₃ displayed high CO conversion (ca. 50% vs 34%) and C₅₊ selectivity (80% vs 76%) compared with their counterparts on unmodified alumina (20 wt % Co, 240 °C, 20 bar, H₂/CO = 2). The enhanced activity of the Ti-modified FT catalyst was ascribed to the suppressed aggregation of cobalt clusters and the better reducibility of Co₃O₄ particles. α -Al₂O₃ supports usually have low specific surface area, and in the direction of improving the specific surface area, Liu et al.⁵¹ demonstrated the preparation of a hierarchical composite material, titania/carbon nanotubes/ α -Al₂O₃ (TiCNTA), consisting of TiO₂-modified multiwall carbon nanotubes (CNTs) decorating a macroscopic host structure of α -Al₂O₃. The composite material exhibited a moderate specific surface area and very open porous structure as well as a strong interaction with the deposited cobalt phase, which in turn significantly improved the dispersion of cobalt nanoparticles and the mass diffusion during the FT reaction. The catalyst supported on TiCNTA showed higher cobalt time yield (6.6 vs 4.2) and C₅₊ selectivity (95% vs 92%) compared with its counterpart on the unmodified support, CNTs/ α -Al₂O₃ (10 wt % Co, 200 °C, 20 bar, H₂/CO = 2). The catalyst also exhibited higher stability compared with other traditional FT synthesis catalysts as a function of time on stream, which was attributed to the medium interaction between the TiO₂ and cobalt NPs. Such stable catalysts will be very useful in application to the biomass-to-liquid process, where some impurities that could be detrimental to the catalyst life cycle are present.

2.2.1.4. Modification with Lanthanum. Lanthanum-modified alumina (γ -Al₂O₃ preimpregnated with lanthanum nitrate before cobalt addition) was reported to influence the dispersion of the metallic cobalt phase, but the extent of its influence depends on its loading.⁵² At low La³⁺ loadings (La/Al atomic ratio \leq 0.026), the presence of La had little effect on the structure of the cobalt catalyst or the CO hydrogenation activity, but the selectivity for long-chain hydrocarbons and olefinic products increased with increasing La content (10 wt % Co, 185 °C, 1 bar, H₂/CO = 3). At higher La³⁺ loadings (La/Al atomic ratio $>$ 0.026), the turnover frequency (TOF) for CO hydrogenation was decreased, probably because of the formation of a La–Co mixed oxide, which eventually decreased the amount of catalytically active cobalt phase. Cai et al.⁵³ demonstrated the preparation of lanthana-modified alumina supports by both impregnation and coprecipitation. They found that cobalt catalysts on the modified supports displayed better reducibility and that more CO molecules were adsorbed on cobalt species, which in turn led to higher activity and lower methane selectivity than with their counterparts on the unmodified support in FT synthesis. Brabant and co-workers systematically studied the effect of the La/Co ratio on the structure of alumina-supported cobalt catalysts in FT synthesis by varying the amount of La (0, 5, 10, 25, 20 wt %) preimpregnated on alumina.⁵⁴ They found that impregnation of the support with 10 wt % lanthanum reduced the concentration of barely reducible cobalt aluminate and also maximized the cobalt metal dispersion. At similar conversions in FT synthesis (10 wt % Co, 220 °C, 20 bar, H₂/CO = 2), lower selectivity for methane and higher selectivity for heavy hydrocarbons were obtained for cobalt catalysts on lanthanum-modified supports. Nevertheless, they displayed higher deactivation rates after several hours of time on stream (60 h) compared with cobalt catalysts on unmodified alumina. They speculated that sintering and the formation of inert carbon phases were possible reasons for catalyst deactivation.

2.2.1.5. Modification with Phosphorus. The chemical modification of γ -Al₂O₃ with phosphorus is widely used in hydrotreatment reactions. The strong affinity between Al and P leads to the formation of AlPO₄, which is known to weaken the metal–support interaction, enhance the reducibility of metal precursors, and stabilize the active phase.⁵⁵ The use of phosphorus-modified γ -Al₂O₃ (P-Al₂O₃) in FT synthesis was reported by Bae and co-workers.⁵⁶ They observed weak interactions between cobalt and the phosphorus-modified alumina surface due to the partial transformation of the γ -Al₂O₃ surface to AlPO₄, which eventually enhanced the reducibility of Co₃O₄ species and inhibited the sintering of metallic cobalt nanocrystallites. The cobalt catalysts on the P-Al₂O₃ support (1–2 wt % P) exhibited higher CO conversion (ca. 40% vs 22%), C₅₊ selectivity (ca. 74% vs 81%), and stability compared with their counterparts on bare Al₂O₃ in a fixed-bed FT synthesis (20 wt % Co, 220 °C, 20 bar, H₂/CO = 2). The observed higher catalytic performance was attributed to the good homogeneous dispersion of cobalt species with high reducibility. The P-Al₂O₃ support also significantly suppressed the migration of cobalt particles to the outer pore mouths by forming thermally stable Co–P oxo species on the Al₂O₃ surface during the FT synthesis reaction, thus helping to enhance the catalytic stability and activity.⁵⁷ Furthermore, the cobalt catalysts supported on phosphorus-modified γ -Al₂O₃ tested in a slurry-phase FT synthesis reaction displayed higher catalytic activity and stability. Modification of

alumina by phosphorus not only suppressed the hydrophilic properties and enhanced the hydrothermal stability of γ -Al₂O₃ but also reduced the deposition of polymeric carbons on the catalytic surface formed during FT synthesis.^{58–60}

2.2.2. Surface Modifications with Carbon/Organic Molecules. The carbon interlayer in carbon-covered alumina-supported cobalt catalysts is known to weaken the detrimental metal–support interactions, thereby avoiding the formation of catalytically inactive species.⁶¹ Jean-Marie et al.⁶² demonstrated that pretreatment of alumina with sorbitol resulted in enhanced cobalt dispersion, but at the cost of the formation of smaller cobalt oxide particles that possess lower reducibility. The cobalt catalysts supported on alumina pretreated with sorbitol exhibited a significant increase in the FT reaction rate (52.5×10^{-3} vs 18.9×10^{-3} mol h⁻¹ g⁻¹; FT reaction conditions: 15 wt % Co, 212 °C, 20 bar, H₂/CO = 2) compared with their counterparts on unmodified alumina. The effect on the catalyst structure and catalytic performance was less pronounced when sorbitol was introduced during impregnation of the cobalt precursor or catalyst post-treatment. These differences were attributed to different mechanisms of decomposition of the cobalt precursor and crystallization of the Co₃O₄ phase in the presence of sorbitol in alumina-supported cobalt catalysts. When β -cyclodextrin was added in very small amounts to the cobalt nitrate impregnating solutions during the preparation of active alumina-supported cobalt catalysts, their catalytic activity in FT synthesis was almost 4 times higher than without the use of β -cyclodextrin.⁶³ The increase in the activity was credited to enhancement of the cobalt dispersion. Carefully choosing the organic molecules or even biopolymers in right amounts with definite functional groups that coat the support material and interact with metal precursors via specific functional groups on the organic moieties would provide a pathway to disperse metals homogeneously on the support surface.

2.2.3. Silylation of γ -Alumina. The presence of a significant concentration of hydroxyl groups on the surface of γ -Al₂O₃ makes it hydrophilic in nature, and furthermore, its surface also exhibits Lewis acidic and Brønsted acidic characteristics.²⁹ Rytter and co-workers chemically modified the surface of γ -Al₂O₃ using organosilanes containing chloro or methoxy active ligands to make the surface of alumina hydrophobic in nature and investigated the effect of surface acidity and hydrophobicity in FT synthesis (20 wt % Co, 210 °C; 20 bar; H₂/CO = 2.1).⁶⁴ They found that preimpregnation with chloro- or methoxyalkylsilanes led to silylation of the γ -Al₂O₃ support, yielding a hydrophobic surface and preventing the formation of highly dispersed cobalt oxide interacting with alumina. However, postimpregnation (i.e., impregnation of silane on the calcined γ -Al₂O₃-supported cobalt catalyst) showed no detectable silylation of the support surface, and the reduction of the catalysts was more challenging because a considerable amount of silanes remained inside the pores. The catalysts supported on γ -Al₂O₃ preimpregnated with silane provided the possibility to improve the selectivity for higher hydrocarbons in FT synthesis. It was hypothesized that silylation of γ -Al₂O₃ yielded a support surface that resembles α -Al₂O₃ and played a role in improving the chain growth probability ($\alpha_{C_{2+}}$).

The effects of alumina surface modifications on the activity and selectivity of cobalt-based FT catalysts are summarized in Table 1. Certainly, the organic or inorganic modifiers play a significant role in altering the direct metal–support interaction,

Table 1. Effects of Surface Modifiers on the Activity and Selectivity of Cobalt Catalysts Supported on Alumina

modifier	catalyst	effects	ref(s)
Mg	Co/Mg/Al ₂ O ₃	<ul style="list-style-type: none"> decreases Co–Al₂O₃ interaction increases Co reduction increases CO hydrogenation for <1 wt % magnesia addition 	42–44
Ba	Co/Ba/Al ₂ O ₃	<ul style="list-style-type: none"> facilitates Co reduction for ≤ 2 wt % BaO addition enhancement in CO conversion and C₅₊ selectivity 	46
Zr	Co/Zr/Al ₂ O ₃	<ul style="list-style-type: none"> decreases Co–Al₂O₃ interaction (0–15 wt %) improves CO conversion and C₅₊ selectivity 	47, 49
Ti	Co/Ti/Al ₂ O ₃	<ul style="list-style-type: none"> alters surface acidic site density increases Co dispersion and reducibility increases CO conversion and C₅₊ selectivity (0.2–0.5 wt %) 	50
La	Co/La/Al ₂ O ₃	<ul style="list-style-type: none"> reduces the amount of cobalt aluminate species and maximizes Co dispersion (10 wt %) increases C₅₊ selectivity 	52–54
P	Co/P/Al ₂ O ₃	<ul style="list-style-type: none"> limits cobalt aluminate formation and aggregation of Co particles (≤ 2 wt %) increases thermal stability and CO conversion 	56, 57
sorbitol (C)	Co/C/Al ₂ O ₃	<ul style="list-style-type: none"> enhances Co dispersion (Co/sorbitol = 6) limits Co reducibility increases CO conversion 	62
silanes (Si)	Co/Si/Al ₂ O ₃	<ul style="list-style-type: none"> modifies the hydrophobicity of the catalyst leads to smaller Co particles improves the chain growth probability ($\alpha_{C_{2+}}$) 	64

which affects the Co reducibility, dispersion, and acid/base properties of the catalyst, thereby influencing the catalytic activity and selectivity.

2.3. Acidity/Basicity of the Oxide Support. The surface acidity/basicity of supported Co catalysts was known to modify the adsorption strengths of CO and H₂ on Co particles, which in turn affected the behavior in FT synthesis.⁶⁵ Surface basicity seemed to increase the electron density of supported Co, thereby strengthening the Co–CO bonds, whereas, surface acidity appeared to decrease the electron density of supported Co, thus strengthening the Co–H bonds. Studies of cobalt supported on γ -alumina possessing different textural properties showed that the support with low acidity displayed higher activity, higher C₅₊ hydrocarbon selectivity, and lower methane selectivity in FT synthesis.⁶⁶ The better catalytic performance of the support with low acidity was attributed to the weaker interaction between metallic cobalt and the support, better reducibility of cobalt, and higher intensity of bridge-mode CO adsorption, which usually favors the dissociation of CO and chain growth of –CH₂– fragments. In order to fully understand the impact of the chemical nature of the support material on the catalytic performance, it is required to decouple this effect from other influential elements such as support texture, extent of Co reduction, and particle size. In this direction, Prieto and co-workers coated a high-surface-area γ -Al₂O₃ (AO) support with a monolayer content of various transition metal (YO_x, TiO_x, TaO_x, WO_x) and lanthanide (SmO_x) oxides (MO_x) to obtain support materials (MO_x@

Table 2. Physicochemical Properties of Oxide Support Materials⁶⁷

support	composition		textural properties					acid/base character η^f (eV)
	metal loading (wt %)	δ^a (atoms/nm ²)	S_{BET}^b (m ² /g)	V_p^c (cm ³ /g)	V_p^d (cm ³ /g _{Al₂O₃)}	PD ^e (cm ³ /g)		
AO	–	–	205	0.52	0.52	8.9	2.49	
SmO _x @AO	18.7 ^g	4.7 ^g	155	0.39	0.50	7.9	2.38	
YO _x @AO	12.1 ^h	4.7 ^h	174	0.44	0.52	7.9	2.46	
TiO _x @AO	6.2	4.2	192	0.46	0.51	8.0	2.50	
TaO _x @AO	23.5	5.3	176	0.34	0.47	7.0	2.53	
WO _x @AO	24.0	5.5	178	0.37	0.53	7.0	2.61	

^aSurface coverage of the alumina surface with M atoms. ^b S_{BET} is the Brunauer–Emmett–Teller specific surface area. ^cTotal specific pore volume. ^dTotal specific pore volume expressed per unit mass of Al₂O₃. ^eAverage pore diameter. ^f η is the spectroscopic parameter, i.e., the energy of the intramolecular charge transfer (IMCT) band of adsorbed alizarin as measured by diffuse-reflectance UV–vis spectroscopy, which gives the Lewis acid/base character. ^gNominal values. ^hDetermined by X-ray fluorescence spectrometry.

Table 3. Physicochemical Properties of CoRu/MO_x@Al₂O₃ Catalysts⁶⁷

catalyst	metal contents (wt %)		textural properties		metal dispersion			
	Co	Ru	S_{BET}^a (m ² /g _{Al₂O₃)}	V_p^b (cm ³ /g _{Al₂O₃)}	$d(\text{Co}_3\text{O}_4)_{\text{XRD}}^c$ (nm)	$d(\text{Co})_{\text{TEM}}^d$ (nm)	H ₂ uptake ^e (μmol/g)	$d(\text{Co})_{\text{H}}^f$ (nm)
CoRu/AO	22.1	0.3	260	0.52	17.3		144	12.5
CoRu/SmO _x @AO	21.8	0.3	292	0.49	12.7	13.6	126	14.1
CoRu/YO _x @AO	21.4	0.2	271	0.45	10.5		162	10.8
CoRu/TiO _x @AO	22.8	0.2	261	0.44	13.3	13.0	132	14.1
CoRu/TaO _x @AO	22.1	0.2	265	0.53	17.4		124	14.6
CoRu/WO _x @AO	19.9	0.2	258	0.47	13.1	16.3	41	39.7

^aBET specific surface area expressed per unit mass of Al₂O₃. ^bTotal specific pore volume expressed per unit mass of Al₂O₃. ^cCo₃O₄ crystallite size after air calcination. ^dSurface-averaged Co particle size distribution as determined by transmission electron microscopy on the reduced and passivated catalysts. ^eTotal hydrogen uptake as determined after in situ reduction. ^fApparent Co average particle size as derived from the total H₂ chemisorption uptake.

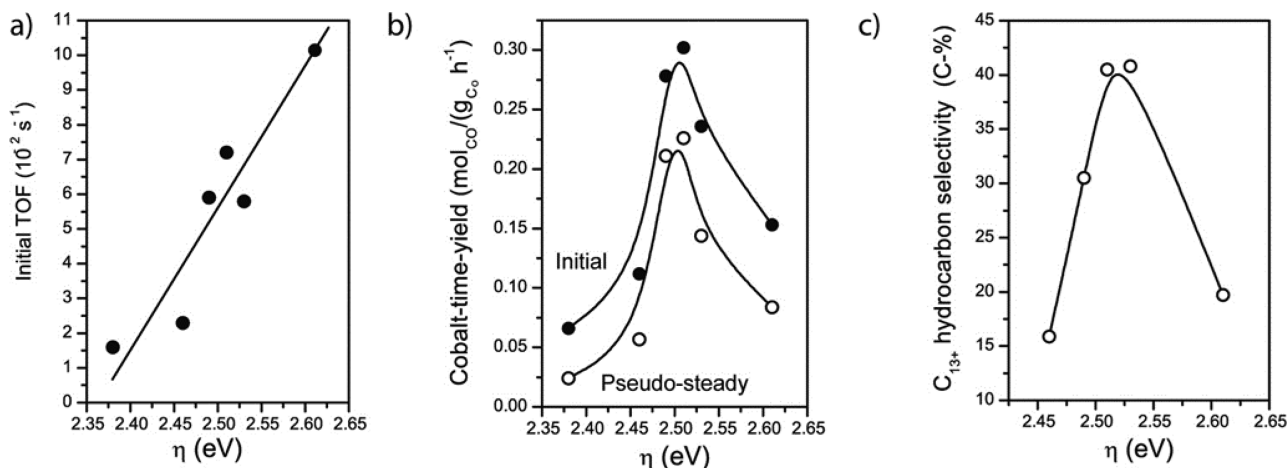


Figure 2. Evolution of (a) the initial turnover frequency (per surface metal atom), (b) the initial (solid symbols) and pseudosteady (open symbols) cobalt time yields, and (c) the C₁₃₊ hydrocarbon selectivity obtained with CoRu/MO_x@AO catalysts as a function of the acid/base character of the support material, represented by the spectroscopic parameter η (see the text). Reaction conditions: 20 wt % Co, $T = 220$ °C, 20 bar, H₂/Co = 2, initial gas hourly space velocity (GHSV) = 6.9 L_{syngas} g_{cat}⁻¹ h⁻¹. Data included in (c) correspond to the steady state after adjustment of the GHSV to achieve a CO conversion level of 40 ± 5% in all cases. Data for CoRu/SmO_x@AO are not included in this panel, as the intrinsically lower activity of this catalyst prevented achieving the mentioned CO conversion level. Lines serve as guides to the eye.⁶⁷

AO) with similar textural properties and vastly different surface acid/base characters (Table 2).⁶⁷ The energy of the lowest intramolecular charge transfer (IMCT)-band of adsorbed 1,2-dihydroxy-9,10-anthracenedione (alizarin) as measured by diffuse-reflectance UV–vis spectroscopy gives the Lewis acid/base character of the material and is denoted as η , the spectroscopic parameter.⁶⁸ Accordingly, SmO_x@AO possesses marked basic character whereas WO_x@AO exhibits strong

Lewis acid character, and the remaining support materials showed intermediate acid/base characters. Of all the supports, only the WO_x@AO support displayed partial reduction at temperatures beyond 830 °C. From temperature-programmed reduction measurements it was concluded that all of the CoRu/MO_x@AO (20 wt % Co, 0.5 wt % Ru) catalysts were quantitatively reduced to metallic Co and Ru after the reduction treatment preceding the catalytic test. Only in the

case of the CoRu/WO_x@AO catalyst was partial coverage of the metal surface by WO_x observed, as concluded from X-ray photoelectron spectroscopy analysis; such a phenomenon is usually termed strong metal–support interaction (SMSI) and is observed for cobalt catalysts supported on partially reducible oxides.⁶⁹ For catalysts supported on relatively reducible oxides such as TaO_x@AO and TiO_x@AO, such effects were minimal. Thus, they ensured that there was no or minimal influence of structural properties such as the support texture, extent of Co reduction, and particle size (in all cases the Co⁰ size was >10 nm, for which no particle size effects on the catalytic activity are expected) (Table 3).

The catalytic test results showed that the initial TOF (per surface Co⁰ atom) increased linearly with increasing acidity of the oxide support (η) (Figure 2a). In contrast, the metal-specific catalytic activity (cobalt time yield) (Figure 2b) and selectivity for C₁₃₊ hydrocarbons (Figure 2c) showed a volcano-type dependence with η , where a maximum was reached for supports of intermediate acid/base character. Furthermore, supports with intermediate acid/base character ($\eta = 2.49\text{--}2.53$ eV) exhibited comparable CH₄ selectivities of ca. 15%. For a more basic oxide support (YO_x@AO), the CH₄ selectivity decreased to ca. 10%, while for a markedly acidic oxide support (WO_x@AO), the CH₄ selectivity increased to ca. 20%. The volcano-type dependence of the metal-specific activity with the nature of the support was attributed to the interplay between the number of surface-exposed cobalt sites and their intrinsic activity. Furthermore, it was inferred from in situ CO FTIR analysis that in the case of basic oxide supports (e.g., SmO_x@AO), few atoms from the basic oxide caused selective blockage on the surface of the metallic Co nanoparticles after reduction, which strengthened CO adsorption and significantly decreased CO hydrogenation; the blocked sites were most likely step edges, which are essential for CO hydrogenation. On the other hand, for stronger Lewis acidic and more reducible oxide supports (e.g., WO_x@AO), unselective site blockage via decoration of metallic Co nanoparticles with oxide patches (termed SMSI) effected a decrease in the number of Co sites available for catalysis. This model study that attempted to delineate structural effects from electronic effects should serve as a guide to develop next-generation catalysts in FT synthesis.

The alumina support is the most extensively studied support in FT synthesis, and the above-discussed surface modification techniques help to improve either the activity or selectivity or both in FT synthesis. For further improvements, studying surface modification with a combination of organic and inorganic species to tune the dispersion and electronic effects of the active phase and performing FT catalytic tests for long duration (time on stream) to establish the catalytic stability would enable their implementation in industry.

3. SILICA-SUPPORTED CATALYSTS

Next to alumina, silica is another widely used support in cobalt-based FT catalysts, as it possesses suitable chemical and physical properties to serve as a perfect support. The nature of hydroxyl groups (silanols) and their concentration and distribution on the silica surface play vital roles in the dispersion of cobalt particles. The high surface area of silica favors moderately high metal dispersion at reasonably high metal loadings. Furthermore, its surface chemistry enables high reducibility of cobalt precursors to the metallic state under favorable conditions. It was reported that the degree of cobalt

reduction was higher on silica compared with other supports, i.e., SiO₂ > TiO₂ > α -Al₂O₃ > γ -Al₂O₃.⁷⁰ Though metal–support interactions are weaker on silica compared with alumina, the preparation and pretreatment conditions must be carefully chosen to avoid the formation of cobalt silicate (Co₂SiO₄), which is inactive in FT synthesis and can be reduced to metallic cobalt only at temperatures above 700–800 °C.^{71–74} The interaction between cobalt and silica that yields cobalt silicate was stronger for the catalysts prepared in basic media than for those prepared in acidic media.⁷⁴ The ease of reduction to metal species on silica supports was found to decrease with decreasing Co₃O₄ particle size.⁷⁵ Noble metal promoters such as Ru, Pt, or Pd were added to cobalt catalysts supported on SiO₂ to improve the reducibility and/or dispersion of cobalt.⁷⁶ In the following sections, the effects of the textural properties, surface modification with different metal oxides or organic molecules, and acid/base properties of silica on the catalytic activity/selectivity in FT synthesis are discussed.

3.1. Effect of Textural Properties of the Silica Support. Cobalt catalysts supported on silica carriers with varying mean pore diameter (4, 6, 10, or 20 nm) tested in the FT reaction (10–12 wt % Co, 202 °C, 20 bar, H₂/CO = 2.5) showed a decrease in CH₄ selectivity with increasing pore diameter.⁷⁷ This study revealed that the product distribution was affected by the porosity of the support, which was due to the variation in the ratio of chemisorbed CO and H₂ on the catalyst surface, i.e., the H₂ uptake decreased with decreasing pore diameter, but the CO uptake remained almost similar irrespective of the pore diameter. Another study found that the metal crystallite size and degree of cobalt reduction increased with increasing pore diameter of the silica support (mean pore diameter of 4, 6, 10, or 15 nm).⁷⁸ Furthermore, the C₅₊ and methane selectivities passed through a maximum and minimum, respectively, in FT synthesis (20 wt % Co, 200–220 °C, 15 bar, H₂/CO = 2) for a cobalt catalyst supported on a silica support having a pore diameter of 10 nm. Song and Li⁷⁹ also reported that cobalt catalysts supported on silica with pore sizes of 6–10 nm exhibited higher activity and higher C₅₊ selectivity in FT synthesis, which were attributed to the moderate Co particle size (ca. 5 nm) and moderate CO adsorption on the catalysts. These studies on silica clearly demonstrated the effect of pore size on controlling the particle size and avoiding diffusional limitations of long-chain hydrocarbons, consequently leading to variation in the uptake of CO and H₂ molecules, which in turn affects the product distribution in FT synthesis.

3.2. Effect of Surface Modifications on the Silica Support. **3.2.1. Surface Modification with Inorganic Oxides.** **3.2.1.1. Modification with Zirconia.** Modification of the silica support with zirconia appeared to positively influence the performance of cobalt catalysts in FT synthesis, particularly when zirconia was added before the impregnation of cobalt.⁸⁰ The addition of zirconia to the silica support during the preparation of the supported cobalt catalyst was found to weaken the Co–SiO₂ interaction by the formation of weaker Co–ZrO₂ interactions that favored the reduction of cobalt at lower temperatures.^{81,82} Consequently, it was found that Zr-modified silica-supported cobalt catalysts exhibited increased steady-state conversion of CO, increased C₅₊ hydrocarbon selectivity, and decreased CH₄ selectivity compared with the unmodified catalysts in FT synthesis (10 wt % Co, 230 °C, 8 bar, H₂/CO = 2, tested for 240 h).⁸² Before the steady-state

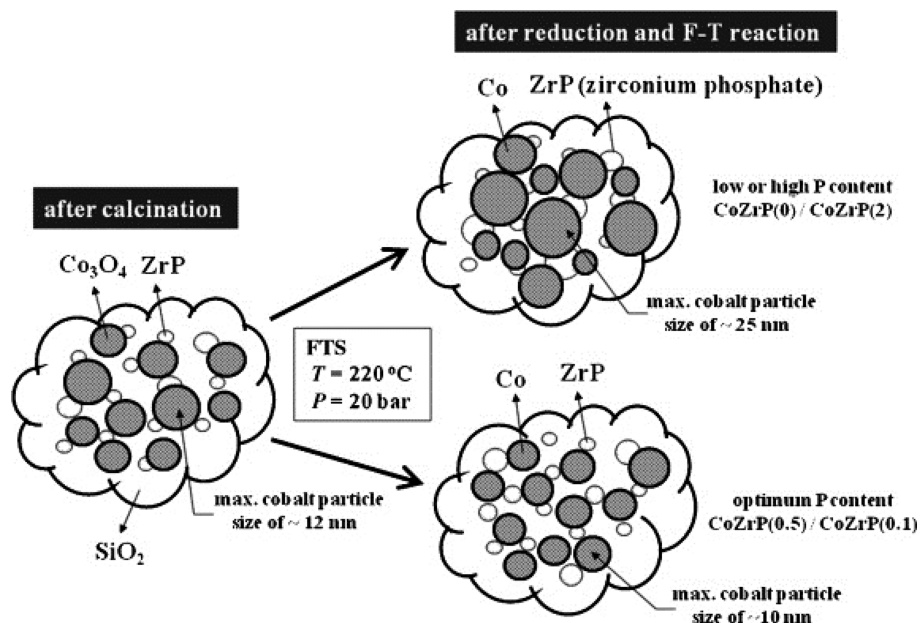


Figure 3. Proposed mechanism of cobalt aggregation in the presence of ZrP particles on the SiO₂ surface. Reprinted with permission from ref 83. Copyright 2011 John Wiley and Sons.

conversion on Zr-modified cobalt catalysts was reached, an initial drop in the activity was observed. Bae and co-workers improved the stability of Zr/SiO₂-based catalysts by adding appropriate amounts of a phosphorus component, yielding zirconium phosphate-modified SiO₂ (ZrP/SiO₂).^{83,84} The cobalt catalysts supported on ZrP/SiO₂ showed a high activity and a low deactivation rate resulting from prevention of aggregation of cobalt particles in FT synthesis, specifically when the P/(Zr + P) molar ratio was in the range of 0.134–0.290.⁸³ The role of ZrP particles on the SiO₂ surface for Co catalysts is schematically depicted in Figure 3. In fact, the enhanced catalytic stability attained using the ZrP/SiO₂ support was attributed to the homogeneous distribution of cobalt clusters, whose mobility was suppressed because the clusters were confined in thermally stable zirconium phosphate matrices on the SiO₂ surface.

Jiang et al.⁸⁵ reported that cobalt catalysts supported on zirconium-modified silica treated using dielectric-barrier discharge (DBD) plasma instead of the conventional thermal calcination step exhibited significantly enhanced cobalt reducibility and increased cobalt dispersion compared with the catalysts on unmodified silica. As a result, plasma-treated cobalt catalysts supported on Zr/SiO₂ displayed higher CO conversions and hydrocarbon chain growth ability in FT synthesis (20 wt % Co, 230 °C, 20 bar, H₂/CO = 2).

3.2.1.2. Modification with Titania. Titania-modified silica showed decreased size of the supported cobalt particles due to the stronger interactions of cobalt oxide species with the titania than with the silica support, consequently increasing the cobalt dispersion without affecting the reducibility of cobalt too much.⁸⁶ The cobalt catalysts supported on titania-modified silica displayed higher activity in FT synthesis (ca. 65% vs 50% CO conversion, 10 wt % Co, 240 °C, 10 bar, H₂/CO = 2) than its counterparts on pristine silica. The high activity of titania-modified catalysts was attributed to the increased cobalt dispersion and increased bridge-type adsorbed CO that was more easily dissociated to carbon and oxygen.⁸⁷

3.2.1.3. Modification with Niobia. In the direction of combining the selectivity-promoting properties of niobia with a highly porous support, den Otter et al.⁸⁸ prepared niobia-modified silica (Nb/Si atomic ratio = 0.02–0.12) by the incipient wetness impregnation method and applied it as a support for cobalt catalysts. They found that at low Nb loadings, Nb existed as small niobia particles or mixed Nb–Si species. However, polymerization occurred at a higher Nb loading (Nb/Si = 0.12) and calcination temperature (900 °C), forming a niobia layer on SiO₂, and eventually 15–20 nm nanocrystals were observed that partially migrated to the exterior surface of the silica grains. Niobia modification of silica did not significantly influence the final cobalt particle size. The cobalt catalysts supported on niobia-modified silica and promoted with platinum yielded an increase in the cobalt-mass-normalized activity in FT synthesis (9–21 wt % Co, 220 °C, 1 or 20 bar, H₂/CO = 2.0) by a factor of 2–3 in the case of amorphous niobia (obtained for a calcination temperature of 550 °C) and by a factor of 3–4 with niobia nanocrystals (obtained for a calcination temperature of 900 °C). The increase in catalytic activity was attributed to an increased number of active sites and an increased cobalt-surface-specific activity. At lower pressure (1 bar), the C₅₊ selectivity in FT synthesis increased by about 50–65% upon niobia modification, but at high pressure (20 bar) it was found to slightly decrease by about 1–4%. The pressure-dependent C₅₊ selectivity of niobia promotion was presumed to be related to a higher intrinsic CO coverage at elevated pressures.

3.2.1.4. Modification with Phosphorus. Cobalt catalysts supported on phosphorus (1 wt %)-modified silica displayed enhanced dispersion of cobalt oxide nanocrystallites and minimized cobalt silicate formation on silica compared with their counterparts on unmodified silica.⁸⁹ Phosphorus-modified silica-supported cobalt catalysts showed higher CO conversions (>60% vs ca. 55%), lower methane selectivity (ca. 7% vs 10%), improved olefin selectivity, and higher C₅₊ selectivity (ca. 87% vs ca. 84%) in comparison with the catalysts supported on unmodified silica (20 wt % Co, 220 °C,

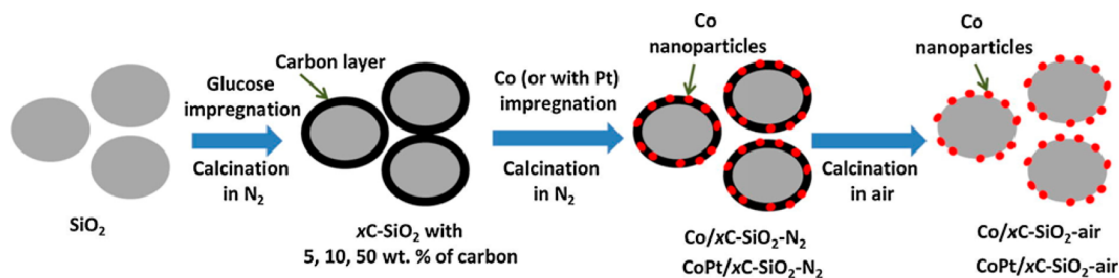


Figure 4. Scheme of the catalyst synthesis on carbon-coated silica. Reprinted with permission from ref 93. Copyright 2016 Elsevier.

20 bar, $H_2/CO = 2$). The enhanced catalytic performance was attributed to the increased cobalt dispersion caused by the altered surface interactions of cobalt with the hydroxyl groups on silica. It was found that the cobalt particles interact with phosphate ions, thereby anchoring Co particles to the support and limiting the cobalt sintering rate.⁹⁰

3.2.1.5. Surface Modification with Carbon/Organic Molecules. Pretreatment of the silica support with an organic solvent such as acetic acid, ethanol, 1-propanol, or 1-butanol before impregnation of cobalt modified the surface properties of silica without altering its textural properties.⁹¹ The pretreatment of silica with acetic acid or alcohols increased the number of isolated silanol (SiOH) groups on the silica surface and hindered the reactivity of H-bonded SiOH moieties with cobalt species, which led to the formation of smaller cobalt particles. Furthermore, pretreatment of silica with organic solvents changed the isoelectric point compared with pristine silica, which resulted in surplus negative charge on the pretreated silica surface that interacts with Co^{2+} ions, yielding improved cobalt dispersion. The cobalt catalysts supported on pretreated supports exhibited higher catalytic activity than their counterparts without pretreatment, which was attributed to the high dispersion and high degree of reduction of the supported cobalt. Among all of the catalysts, the cobalt catalyst supported on silica pretreated with acetic acid showed the highest CO conversion and the lowest CH_4 selectivity. Chen et al.⁹² demonstrated modification of the silica surface with ethylene glycol, which enhanced the cobalt dispersion. Cheng and co-workers showed that coating the silica support with carbon (5, 10, or 50 wt %) by impregnation of silica with glucose and subsequent calcination (450 °C, 4 h) led to stabilization of highly dispersed cobalt oxide particles with partial cobalt reduction during the thermal treatment in an inert (N_2) atmosphere (Figure 4).⁹³ The sizes of the cobalt oxide nanoparticles on the carbon (50 wt %)-coated silica-supported catalysts (5–8 nm) were considerably smaller than those of the counterparts without carbon coating (10–15 nm). The FT activity of catalysts prepared with higher carbon content, calcined in nitrogen followed by air, and containing smaller cobalt particles was almost 2 times higher compared with the reference uncoated catalysts (15 wt % Co, 220 °C, 20 bar, $H_2/CO = 2$). Furthermore, higher selectivity for long-chain hydrocarbons and low methane selectivity were also observed. It was speculated that the selectivity effects might arise from different factors such as the formation of cobalt nanoparticles with fewer defects in the presence of carbon, selective poisoning of methanation sites by residual carbon, higher cobalt reducibility, and lower content of cationic cobalt silicate species.

The effects of surface modifications of the silica support on the activity and selectivity of cobalt-based FT catalysts are

summarized in Table 4. As observed in the case of alumina, silica surface modifiers exhibit a significant role in altering the

Table 4. Effects of Surface Modifiers on the Activity and Selectivity of Cobalt Catalysts Supported on Silica

modifier	catalyst	effects	ref(s)
Zr	Co/Zr/SiO ₂	<ul style="list-style-type: none"> limits Co–SiO₂ interactions favors weak Co–Zr interactions increases Co reducibility enhances CO conversion and C₅₊ selectivity (2–10 wt %) 	80–82
ZrP	Co/ZrP/SiO ₂	<ul style="list-style-type: none"> prevents Co particle aggregation increases steady-state CO conversion and catalytic stability (with P/(Zr + P) molar ratio of 0.029–0.134) 	83, 84
Ti	Co/Ti/SiO ₂	<ul style="list-style-type: none"> enhances Co dispersion modifies Co reducibility increases CO conversion (2–10 wt %) 	86
Nb	Co/Nb/SiO ₂	<ul style="list-style-type: none"> increases the number of active sites improves the cobalt-mass-normalized activity (Nb/Si atomic ratio = 0.02–0.12) 	88
P	Co/P/SiO ₂	<ul style="list-style-type: none"> limits Co sintering improves catalyst stability increases Co conversion (≤1 wt %) 	89, 90
organic molecules (C)	Co/C/SiO ₂	<ul style="list-style-type: none"> facilitates Co dispersion and reducibility increases CO conversion and C₅₊ selectivity (≤10 wt %) 	91, 93

detrimental cobalt silicate formation as well as the Co reducibility and dispersion, consequently influencing the catalytic activity and selectivity.

4. MESOPOROUS SILICA-SUPPORTED CATALYSTS

Periodic mesoporous silicate materials prepared using surfactant-templated hydrothermal synthetic procedures are evolving as promising support materials for catalytic metal precursors because of their characteristics such as narrow pore size distributions, high surface areas and pore volumes, and controllable acid/base properties.⁹⁴ These properties usually lead to higher metal dispersion in the supported metal catalysis, which eventually leads to an increase in the surface density of active metal sites. MCM-41 and SBA-15 are the two most common types of periodic silicas used as catalytic carriers. MCM-41 exhibits a unidimensional and hexagonal arrangement of mesopores with dimensions in the range of 20

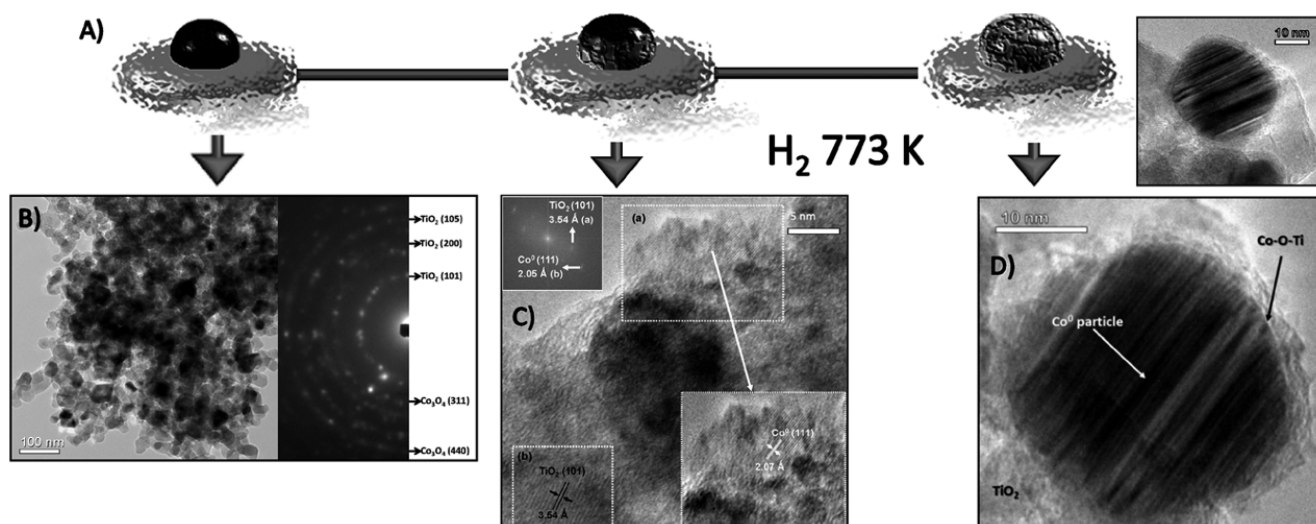


Figure 5. (A) Scheme of the SMSI encapsulation effect. (B) TEM image and electron diffraction pattern for a Co10-ct sample (calcined Co catalyst on TiO₂). (C) HRTEM image and its Fourier transform for Co10-ctr (reduced Co catalyst on TiO₂): (a) cobalt nanoparticles over TiO₂ (the reconstructed image corresponding to fcc-Co(111) is shown at the bottom right); (b) lattice fringes for TiO₂. (D) HRTEM images of a Co10-ctr sample depicting the decoration of Co metal nanoparticles with an amorphous TiO₂ layer. Republished with permission from ref 103. Copyright 2011 Royal Society of Chemistry.

Å to greater than 100 Å that are tuned by varying the alkyl chain length of the surfactant template and the synthesis conditions.⁹⁵ SBA-15 contains two-dimensional hexagonally ordered mesopores with pore sizes of up to 300 Å that are obtained using a variety of poly(alkylene oxide) triblock copolymers and by the addition of cosolvent organic molecules.⁹⁶ SBA-15 mesoporous material also displays good thermal, mechanical, and chemical resistance properties, which are essential characteristics for a material to be used as a catalyst carrier. Conventional catalyst carriers usually exhibit broad pore size distributions, making it difficult to draw unambiguous conclusions about the influence of catalyst pore sizes on FT reaction rates and selectivities. The tailored pore size distributions in periodic mesoporous silicas such as MCM-41 and SBA-15 make these materials suitable to be used as model catalytic supports in evaluating the effects of pore size in FT synthesis.^{97,98} The small, narrow pores present in MCM-41 materials lead to the formation of smaller cobalt oxide particles that are quite difficult to reduce. Consequently, smaller Co species located in the narrow-pore silicas displayed much lower activities in FT synthesis with higher methane selectivities than larger cobalt particles in the wide-pore supports. SBA-15-type periodic mesoporous silicas with a very narrow pore size distribution (ca. 9.1 nm) exhibited higher cobalt dispersion over commercial silicas with a broad pore size distribution (ca. 33 nm), even at higher cobalt surface densities (5–15 Co/nm²) that resulted in higher cobalt time yields in FT synthesis.⁹⁹ At high cobalt surface densities, commercial silicas exhibited sintering of Co₃O₄, but on SBA-15, high cobalt dispersion remained intact. Despite the advantages offered by these mesoporous materials, there are some limitations. For example, aqueous impregnation of MCM-41 with the cobalt precursor followed by drying leads to the loss of long-range ordering of the hexagonal mesoporous structure and to a significant decrease in the surface area and pore volume.⁹ However, the SBA-15 mesoporous structure seemed to remain almost intact after introduction of significant amounts of cobalt (up to 20%). The difference in structural stability of these two

mesoporous silicas was attributed to the different pore wall thicknesses in MCM-41 and SBA-15. Further work is needed to establish its long-term stability in FT synthesis by its surface modification with inorganic and/or organic species.

5. TITANIA-SUPPORTED CATALYSTS

Titania usually exhibits relatively low specific surface areas (ca. 60 and ca. 10 m² g⁻¹ for the anatase and rutile allotropes, respectively) compared with alumina and silica supports. This limits the cobalt loading in FT catalyst formulations to maintain high metal dispersion. Furthermore, TiO₂ is known to exhibit the SMSI effect, as first reported by Touster et al.¹⁰⁰ This type of SMSI is different from the other types of solid-state metal–support interactions, where the metal precursor chemically reacts with the support to form catalytically inactive (refractory) species. SMSIs altered the H₂ and CO chemisorption properties of group VIII metals when these elements were supported on titania. The SMSI effect could be explained via electron transfer from the support to the metal (electronic factor)¹⁰¹ or by the formation of intermetallic phases.¹⁰² In the case of titania (anatase)-supported cobalt catalysts (Co/TiO₂), during reduction of the cobalt oxide particles, partial reduction of the TiO₂ support takes place, giving rise to TiO_x suboxide species ($x < 2$) that migrate onto the surface of cobalt particles and encapsulate or decorate the metallic cobalt particles. The evolution of the SMSI decoration effect, as illustrated by high-resolution transmission electron microscopy (HRTEM) images, is depicted in Figure 5.^{103,104} Under similar conditions, the catalytic performance of Co/TiO₂ for CO hydrogenation was very low (<5%) compared with Co/SiO₂ (ca. 30%) (10 wt % Co, 230 °C, 40 bar, H₂/CO = 2).¹⁰³ Its low activity in FT synthesis was attributed to blocking of the active Co⁰ phase caused by the migration of Ti species during the reduction step. Despite the active-phase blocking effect, TiO₂-supported cobalt catalysts were reported to display high selectivity for long-chain hydrocarbons (C₅₊) compared with their counterparts on alumina.^{105–107}

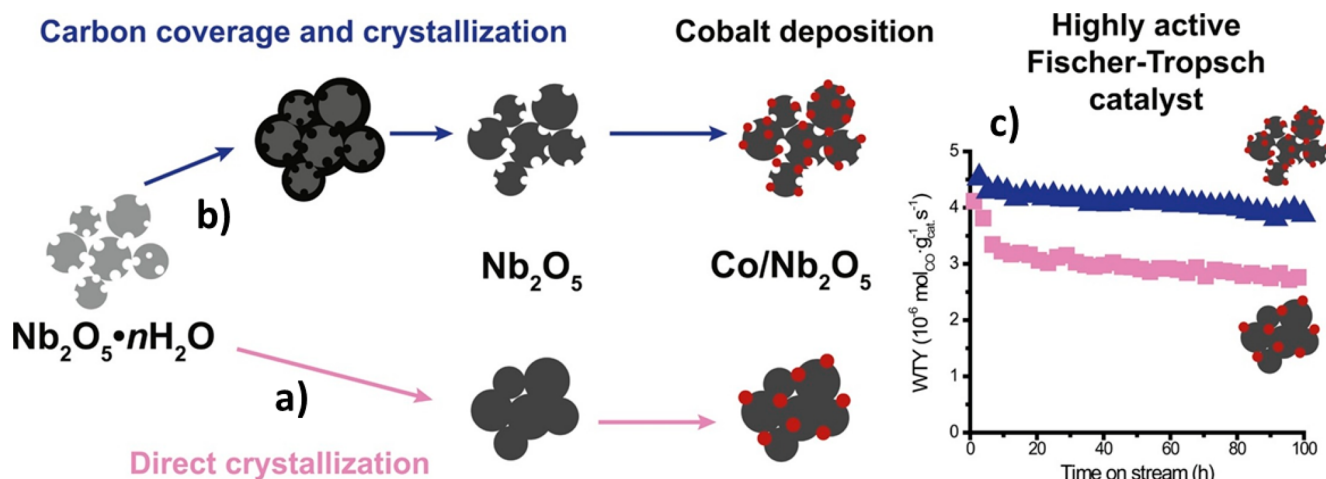


Figure 6. (a) Dense Nb₂O₅ obtained by calcination at 600 °C. (b) Porous Nb₂O₅-MC obtained by sucrose impregnation, pyrolysis, and calcination. (c) Activity of Nb₂O₅-supported cobalt catalysts (blue triangles, 10 wt % Co/Nb₂O₅-MC catalyst; pink squares, 6 wt % Co/Nb₂O₅) normalized by catalyst mass (weight time yield, WTY). Reproduced from ref 116, published by Elsevier.

Iglesia et al.⁸⁷ demonstrated that the addition of Ru to Co/TiO₂ (Ru/Co atomic ratio < 0.008) increased both the turnover rate and C₅₊ selectivity in FT synthesis. The improved activity was attributed to the higher cobalt site density, enhanced reduction of cobalt oxide species, and inhibition of carbon deposition reactions resulting from the intimate mixing between Co and Ru. Moreover, Eschemann and co-workers demonstrated that the addition of noble metals such as Pt, Ru, Re, and Ag resulted in decreased reduction temperatures of cobalt oxide supported on TiO₂, eventually leading to increases in the CO hydrogenation activity in FT synthesis.¹⁰⁸ Cobalt catalysts supported on TiO₂ whose surface was modified with boron (0.05 wt %) showed slightly increased syngas conversion only at higher space times (>0.33 h g_{cat}/NL), but the product selectivity in FT synthesis remained the same.¹⁰⁹ Conventionally, FT catalysts are reduced under a hydrogen atmosphere before testing in FT synthesis, but Jalama and co-workers demonstrated that the activation of a titania-supported cobalt catalyst with syngas yielded a catalyst that showed higher stability, higher activity, and better product selectivity (C₅₊ and C₃-C₅ olefin to paraffin ratio) in FT synthesis than its counterpart activated with hydrogen gas (20 wt % Co, 220 °C, 20 bar, H₂/CO = 2).¹¹⁰ The enhanced performance of the catalyst activated with syngas was attributed to increased active-site density resulting from improved catalyst reduction and dispersion. Furthermore, it was proposed that CO prevented the formation of inactive Co species that interact strongly with the support.¹¹¹

6. NIOBIA-SUPPORTED CATALYSTS

Niobium oxide (Nb₂O₅) is known to be a water-tolerant solid acid catalyst, and furthermore, its calcination or pretreatment temperature plays a key role in determining its acidic and catalytic properties.¹¹² As observed in the case of titania, Nb₂O₅ as a reducible catalytic carrier also exhibits SMSI that is more efficient than that of titania.¹¹³ Furthermore, Nb₂O₅-supported cobalt catalysts were reported to display high selectivity for heavy hydrocarbons (C₅₊ selectivity ca. 49%) in FT synthesis at atmospheric pressure (5 wt % Co, 260 °C, 1 bar, H₂/Co = 2) compared with its counterparts supported on alumina (C₅₊ selectivity ca. 11%).¹¹³ The chain-growth-promoting effect of Nb₂O₅ as a support material was attributed

to partial reduction of the support and SMSI.¹¹⁴ The C₅₊ selectivity of Nb₂O₅-supported cobalt catalysts was also reported to be very high, i.e., up to 90 wt % C₅₊ under industrial working conditions (5–20 wt % Co, 220 °C, 20 bar, H₂/Co = 2).¹¹ It was found that 5 wt % Co/Nb₂O₅ (operated at ca. 250 °C) and 25 wt % Co/γ-Al₂O₃ (operated at ca. 210 °C) have identical activities per unit mass of catalyst at similar C₅₊ selectivities (ca. 80 wt %), and consequently, the activity per unit mass of cobalt was 4 times higher for the Nb₂O₅-supported catalyst. However, it should be noted that the catalysts supported on Nb₂O₅ required a higher operating reaction temperature, i.e., 40 °C higher than their Al₂O₃ counterparts. Furthermore, when 6 wt % Co/Nb₂O₅ catalysts were promoted with platinum (Co/Pt atomic ratio ca. 140), the cobalt-mass-normalized activity in FT synthesis (6 wt % Co, 220 °C, 20 bar, H₂/Co = 2) increased by a factor of 2.4, while the high C₅₊ selectivity (85 wt %) was maintained.¹¹⁵ The increase in the activity of Pt-promoted Co/Nb₂O₅ catalysts was tentatively credited to the more efficient promotion of Co by NbO_x (x < 2.5) species resulting from the reduction of Nb₂O₅ and to an increase in the number of catalytically active sites by a factor of 1.7 as evidenced from CO chemisorption experiments.¹¹⁵

Multiple impregnations on crystalline Nb₂O₅ to achieve cobalt loadings higher than 6 wt % usually led to larger cobalt particles, reduced interaction of cobalt with Nb₂O₅, and subsequently decreased activity per unit mass of cobalt and C₅₊ selectivity.¹¹⁶ Preparation of Nb₂O₅-supported catalysts with higher cobalt loadings and dispersion is very challenging because of the very low specific pore structure of crystalline Nb₂O₅ compared with commercial FT supports such as silica and alumina. In the direction of circumventing this challenge, Hernández Mejía and co-workers reported the preparation of mesoporous crystalline Nb₂O₅ (Nb₂O₅-MC) with tailored porosity by carbon deposition via sucrose pyrolysis (Figure 6a,b).¹¹⁶ Deposition of carbon derived from sucrose during the crystallization step restricted the collapse of the Nb₂O₅·nH₂O porous structure. The larger specific pore volume of Nb₂O₅-MC enabled those authors to achieve higher cobalt loading (10 wt %) in a single impregnation step. The higher cobalt loadings in Co/Nb₂O₅-MC led to an increased catalyst-mass-normalized activity in FT synthesis (Figure 6c). The porosity

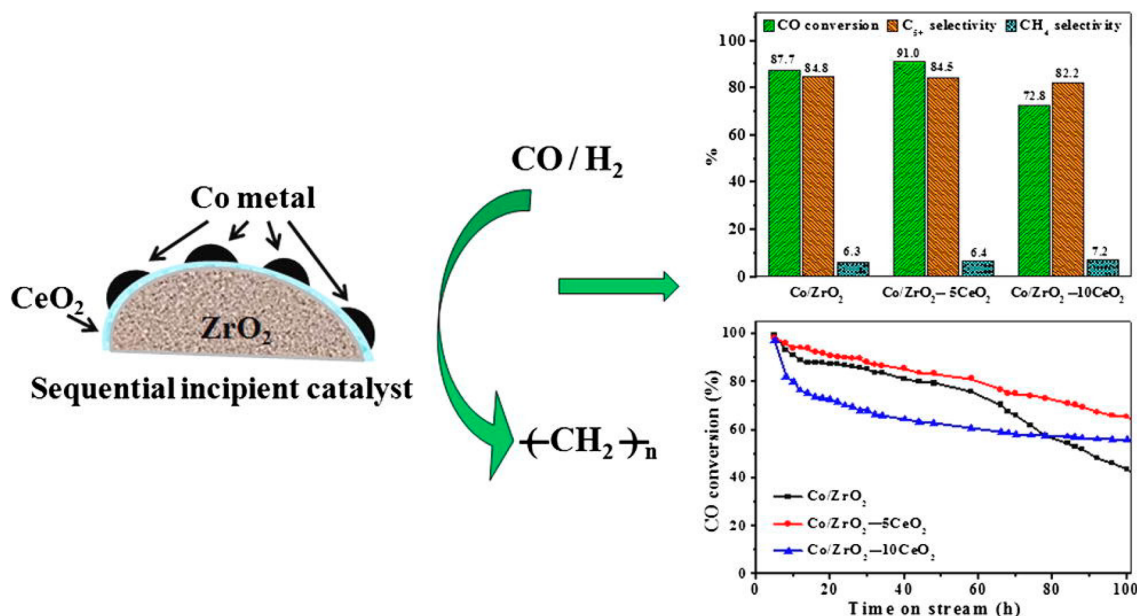


Figure 7. Cobalt catalysts supported on ZrO₂ and CeO₂-modified ZrO₂ prepared by the sequential incipient impregnation method and their activity and hydrocarbon selectivity results in FT synthesis. Reprinted with permission from ref 122. Copyright 2016 Elsevier.

of Nb₂O₅-MC prevented sintering of the cobalt particles, resulting in relatively stable catalysts.

7. ZIRCONIA-SUPPORTED CATALYSTS

Zirconia is drawing considerable interest on account of its potential use as a catalyst support due to its chemical inertness, redox properties, and acid/base characteristics.¹¹⁷ A ZrO₂ support was found to promote the formation of amorphous hexagonal metallic cobalt during the reduction step in the catalyst preparation. Furthermore, the direct reduction of cobalt nitrate leads to weaker metal–support interactions than observed in the case of calcined catalysts, consequently increasing the quantity of amorphous or poorly crystalline hexagonal metallic cobalt, which has been proved to be more active than cubic cobalt in FT synthesis.¹¹⁸ ZrO₂-supported cobalt catalysts exhibited better Co reducibility and capability of hydrogen adsorption via a spillover mechanism compared with conventional alumina-supported Co catalysts.¹⁰ These properties resulted in a higher catalytic activity (CO conversion, 45% vs 36%) and an increase in the chain growth probability (C₅₊, 57% vs 48%) in the FT synthesis (10.5 wt % Co, 230 °C, 5 bar, H₂/CO = 9). Commercially available micrometer-sized (50 m²/g)- and nanosized (78 m²/g) ZrO₂ usually possess medium specific surface areas. Panpranot et al. demonstrated the preparation of nanocrystalline ZrO₂ by decomposition of zirconium tetra *n*-propoxide in 1,4-butanediol that exhibited a high specific surface area (170 m²/g).¹¹⁹ Glycothermal-derived nanocrystalline ZrO₂-supported cobalt catalysts displayed higher activity, higher selectivity for C₄–C₆ hydrocarbons, and lower selectivity for methane compared with the commercial micrometer-sized and nanosized ZrO₂ in FT synthesis (8 wt % Co, 220 °C, 1 bar, H₂/CO = 10). The better catalytic performance was attributed to higher cobalt dispersion and the presence of highly stable and pure tetragonal-phase zirconia. Liu and co-workers reported the preparation of mesoporous zirconia with different pore sizes possessing the characteristics of both zirconia and mesoporous material.¹²⁰ The cobalt catalysts supported on

mesoporous zirconia exhibited enhanced average crystalline sizes and degrees of reduction of Co₃O₄, and in addition, the strength of the Co–ZrO₂ interaction decreased with increasing pore diameter (5.3 to 6.1 to 8.9 nm). The catalytic test results displayed enhanced activity, higher C₅₊ selectivity, and decreased methane selectivity in FT synthesis with increasing pore size. The enhanced catalytic activity was attributed to the improved mass transfer derived from the large pore size and narrow pore size distribution of the mesoporous zirconia, and the improved C₅₊ selectivity was derived from the use of zirconia as the catalytic support.

7.1. Effect of Surface Modifications on the Zirconia Support. Cobalt catalysts supported on lanthanum (10 mol %)-modified ZrO₂ (La-ZrO₂) were prepared by mechanical mixing and exhibited higher cobalt dispersion and enhanced cobalt reducibility, consequently yielding higher catalytic activity in FT synthesis than those supported on the pure oxides (La₂O₃ and ZrO₂) and other La-modified ZrO₂.¹²¹ The other La-ZrO₂ was prepared by coprecipitation and impregnation, causing intimate contact between La₂O₃ and ZrO₂ and incorporated La atoms in ZrO₂, which in turn induced stronger interactions between cobalt and the support and lowered the catalytic activity in FT synthesis. Another study showed that cobalt catalysts supported on ceria (5 wt %)-modified ZrO₂ (5CeO₂-ZrO₂) prepared by the sequential incipient impregnation method (Figure 7) also had enhanced cobalt oxide dispersion, cobalt reducibility, and number of cobalt active sites compared with those supported on unmodified ZrO₂.¹²² Subsequently, the Co/5CeO₂-ZrO₂ catalysts displayed higher activity in FT synthesis (15 wt % Co, 220 °C, 20 bar, H₂/CO = 2). The better catalytic performance was attributed to the cyclic redox process of CeO₂, which can effectively inhibit the oxidation of Co⁰ to Co²⁺. However, the C₅₊ selectivities on the CeO₂-modified and unmodified catalysts were the same (ca. 85%). Addition of an excessive amount of CeO₂ (10 wt %) resulted in decreased activity in FT synthesis, which was ascribed to rapid catalyst deactivation in the initial stages of the

reaction caused by diffusional limitations of the reactants/products (liquid waxes) in the smaller pores.

8. ZEOLITE-SUPPORTED CATALYSTS

Zeolites are crystalline microporous aluminosilicates with frameworks assembled from tetrahedral units having a cation such as Si or Al at the center and oxygen atoms at the corners. The utility of zeolites as catalytic supports is generating considerable interest in heterogeneous catalysis because of their shape-selective characteristics, well-defined pore structure, high porosity, and acidic nature.¹²³ Zeolites can be synthesized with various structures and pore dimensions. To simplify the notation of diverse zeolite structures, a structure-type code was introduced by IUPAC (e.g., a zeolite containing 10-membered-ring pore systems that form a three-dimensional network is named ZSM-5). The acidic nature of the zeolites can be tuned by varying the silica-to-alumina ratio. The surface acidic nature of zeolites tends to display FT product selectivity for lighter and branched hydrocarbons and even aromatics, thus deviating with respect to Anderson–Schulz–Flory (ASF) distributions for cobalt catalysts in FT synthesis.^{124–126}

Furthermore, zeolite-supported cobalt catalysts serve as a bifunctional catalyst composition in which metallic cobalt drives the growth of long-chain hydrocarbons, which undergo further reaction at the zeolite acid sites to yield more-branched hydrocarbons with limited chain length and aromatics.¹²⁷ ZSM-5-supported cobalt catalysts showed good selectivity for gasoline-type products in FT synthesis due to the characteristic properties of ZSM-5 such as surface acidity (which promotes secondary reactions like oligomerization, isomerization, cracking, and aromatization), pore structure (which provides selectivity for limiting the hydrocarbon chain growth length), resistance to coke formation, and stability under FT reaction conditions.^{13,128–130} It was found that the presence of secondary porosity, cages, and three-dimensional channel systems with large micropores in zeolites (e.g., USY) mediates higher accessibility of reactants/products to the catalytic sites, thereby favoring the formation of long hydrocarbons.¹³¹

β -Zeolite is classified as a large-pore zeolite that possesses two types of 12-membered pores and a three-dimensional structure with both micro- and mesopores.¹³² The cobalt catalysts supported on mesostructured β -zeolite showed higher activity and lower methane selectivity in FT synthesis than those supported on β -zeolite, which was ascribed to the better reducibility of larger cobalt particles that exhibit weaker interactions with the support.¹³³ In another study, cobalt catalysts (7.5, 10, 15, or 20 wt %) supported on nanometric β -zeolite composed of β -zeolite crystallites of nanometric dimensions and a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of about 50 were reported.¹³⁴ These catalysts exhibited bifunctional character with a maximum activity at a specific concentration of cobalt (15 wt %) and Co_3O_4 particle size (ca. 9). However, the selectivity for isoparaffins from isomerization reactions was strongly influenced by the support's acid functionality. A recent study found that the catalytic activity and product selectivity of cobalt supported on β -zeolite in FT synthesis strongly depends on the catalyst preparation method, i.e., impregnation (IM), incipient wetness (IW), physical mixing (PM), or precipitation (PT).¹³⁵ The activities of the catalysts were ranked in the following order: IW > IM > PT > PM. The selectivities for isoparaffins were placed in the following order: PT > PM > IM > IW.

Generally, most types of zeolite material consist of pore sizes and cavities in the microporous regime, which hampers the transport of reactants and products during FT synthesis. In the direction of overcoming this limitation, hierarchical zeolites were developed that combine the advantages of mesoporous materials and zeolite crystals.¹³⁶ The cobalt catalysts supported on hierarchical zeolite Y displayed increased activity and C_{5-11} selectivity in FT synthesis (10 wt % Co, 260 °C, 10 bar, $\text{H}_2/\text{CO} = 2$) compared with their counterparts supported on pristine zeolite Y (Figure 8).¹³⁷ The better catalytic perform-

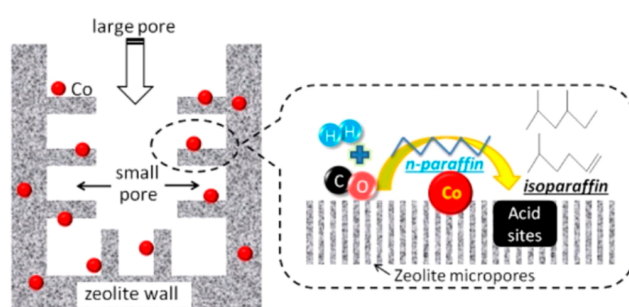


Figure 8. Schematic depiction of the hierarchical zeolite and its use as a catalytic support in FT synthesis. Republished with permission from ref 137. Copyright 2015 Elsevier.

ance was attributed to the optimized hydrocracking and isomerization function derived from the hierarchical zeolite Y with the strong Brønsted acid/Lewis acid (B/L) ratio and effective textural properties. In a different study, a hierarchically spherical Co-based zeolite catalyst with an aggregated nanorod structure prepared by in situ self-assembly crystallization of Co/SBA-15 also exhibited improved activity and enhanced the isoparaffin selectivity in FT synthesis.¹³⁸

Lu et al.¹³⁹ demonstrated a new method of preparing H-USY-zeolite-supported nanocobalt bifunctional catalysts by a physical sputtering process (Figure 9). In this process, the

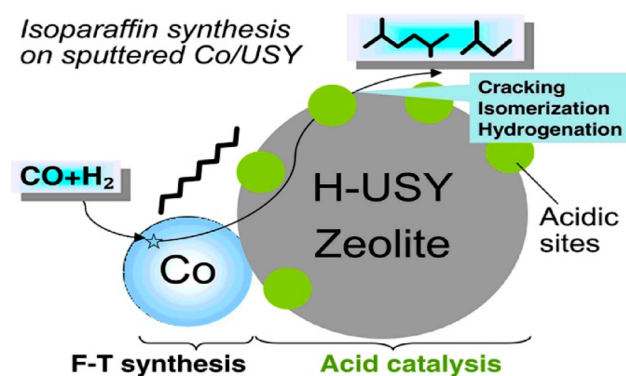


Figure 9. Schematic of the tandem reaction including FT synthesis and acid catalysis achieved on the sputtered Co/USY bifunctional catalyst. Republished with permission from ref 139. Copyright 2015 Elsevier.

particles of sputtered cobalt exist at the nanolevel and were well-dispersed on the acidic USY zeolite. The nanocobalt loaded on the support was easily activated in a low-temperature (260 °C) hydrogen reduction atmosphere. The sputtered Co metallic nanoparticles wedged on zeolite form only weak physical interactions between the metal and the support, thus leading to a facile metal reduction process in this

case, in contrast to the strong interaction between the impregnated metal and the support. Consequently, the sputtered bifunctional Co/USY catalyst exhibited much higher CO conversion and higher isoparaffin selectivity than the conventional impregnated catalyst (7 wt % Co, 260 °C, 10 bar, H₂/CO ≈ 2). This method might be useful in heterogeneous catalysis where detrimental SMSI effects are observed.

9. CERIA-SUPPORTED CATALYSTS

CeO₂ is an interesting choice for use as a catalyst support, considering its redox properties, reducible nature, and surface affinity for both CO and H₂ molecules. Furthermore, it can also function as a structural and electronic promoter in heterogeneous catalysis.¹⁴⁰ Usually, under typical FT synthesis conditions using cobalt catalysts supported on conventional supports (5 wt % Co, 220 °C, 20 bar, H₂/CO = 2), the formation of oxygenates is often neglected. However, it was observed that CeO₂-supported cobalt (ca. 5 wt %) catalysts prepared by the incipient wetness impregnation method exhibited high selectivities for oxygenates (ca. 7%) and C₂–C₄ hydrocarbons (ca. 38%) and low selectivity for C₅₊ hydrocarbons (44%) in the FT synthesis process (5 wt % Co, 250 °C, 20 bar, H₂/CO = 2).¹² Though it might be argued that the temperature was slightly higher (250 vs ca. 220 °C) in the above study, the selectivity for oxygenates was still remarkable. Gnanamani et al.¹⁴¹ proposed a likely pathway for alcohol formation based on Co–CeO₂ interfacial catalysis, in which FT reaction intermediates at the cobalt metal sites diffuse from the cobalt metal particles across the Co–CeO₂ junction to react with active bridging OH groups on partially reduced ceria (Ce³⁺) groups to form oxygenates. Furthermore, it was found that CeO₂-supported Co (5, 15, or 25 wt %) catalysts promoted with Pt exhibit higher selectivities for linear alcohols (14.8%) at lower cobalt loadings (5 wt % Co, 220 °C, 20 bar, H₂/CO = 2). Increasing the cobalt content led to a decrease in the selectivity for oxygenates (8.9% and 4.5% for 15 and 25 wt % Co loading, respectively) because the on-top surface area of cobalt increased at the expense of the interfacial area at the Co–CeO₂ junction.

10. CARBON-SUPPORTED CATALYSTS

Carbon-based materials such as carbon nanofibers (CNFs), carbon nanotubes (CNTs), carbon spheres (CSs), and mesoporous carbons have drawn interest as catalyst supports in FT synthesis because of their diverse merits such as high surface areas with diverse pore structures, high purity, high mechanical strength and thermal stability, good electrical conductivity, resistance to acidic and basic conditions, low cost, good recycling characteristics, low density, and superior chemical inertness under FT reaction conditions.^{142,143} In general, carbon is classified into two types: amorphous carbon, which comprises activated and templated carbon, and crystalline carbon, which comprises graphite, fullerenes, CNTs, carbide, and diamond. The chemical inertness of pristine carbon materials helps in circumventing the strong detrimental interactions between the active metal and the support, thereby enabling studies of the catalytic behavior of the metal with respect to its structure, size, and shape in the absence of support effects. The articles reporting on cobalt-based carbon catalysts prior to the year 2015 were summarized in two reviews. In the first one, progress on carbon support structure effects, the influence of surface modification of the

carbon pore structure by O and N doping, and the effects of cobalt intrinsic properties and promoters on carbon-supported cobalt catalysts for FT synthesis were presented.¹⁴³ In the second one, the features and surface properties of several shaped carbon materials, including carbon black, activated carbon, CNTs, CNFs, CSs, ordered mesoporous carbon, graphene, and diamond and their utility as catalytic supports in FT synthesis and higher alcohol synthesis were discussed and summarized.¹⁴⁴ In the following paragraphs, only the important findings and the articles from 2015 to the present are discussed.

The effect of Co particle size on traditional oxides in FT synthesis is still not very well established because of the varying degrees of interaction between the support and the metal. In the absence of strong interactions between Co and CNFs, it was found that cobalt particle sizes of 6–8 nm are preferred for obtaining both higher TOF and C₅₊ selectivity.^{145,146} The lower TOF of <6 nm Co particles was attributed to blocking of edge/corner sites and lower intrinsic activity at the small terraces.¹⁴⁷ Though the reducibility of cobalt oxide species is improved on carbon material supports, the preparation of highly dispersed and stable cobalt catalysts still requires at least intermediate interactions between the support and the Co precursor. Modification of surface chemical properties of carbon materials by O or N doping was found to be essential to increase their ability to anchor cobalt species. Oxygen-containing groups (e.g., carboxylic acid/anhydride, lactone, and phenolic hydroxyl) were introduced on the surface of carbon by treatment with nitric acid, thus increasing the surface acidity and adsorption ability to anchor Co particles.¹⁴⁸ The acid-treated CNTs yielded a decrease in cobalt particle size and an increase in cobalt dispersion, consequently positively affecting the activity in FT synthesis.^{149,150} Nitrogen functional groups were introduced on the surface of carbon by treatment with ammonia or aniline, thereby altering the defect and electronic structure of the support.¹⁴² The N-functionalized CNTs as a cobalt catalyst support exhibited increased metal–support interactions, Co dispersion, and reducibility of cobalt oxides, and these observed properties led to a distinct shift to lower-molecular-weight hydrocarbons and provided excellent stability to the N-functionalized CNT-supported cobalt catalyst.¹⁵¹ It was also found that N doping of CSs not only improved the cobalt particle distribution but also influenced the reduction ability of carbon. Thus, N-doped CS-supported Co catalysts autoreduced by the carbon support showed higher CO conversion than that obtained by H₂ reduction.¹⁵² CNTs are distinguished from other carbon materials by their graphite layers and tubular morphology, which impart remarkable electronic and mechanical properties.¹⁵³ The channels in CNTs restricted cobalt particle growth inside the tubes (confinement effect), and as a result, particles inside the tubes were smaller than the particles attached to the outer surfaces.¹⁵⁴ Furthermore, it was found that the sintering rate of the particles located on the outer surface of the CNTs was significantly higher than that of particles on the inner surface because of the electron deficiency of the inner walls of the CNTs and the stronger interaction between the cobalt oxides and the support. The confinement of the reactant and reaction intermediates inside the channels increases the contact time of these reaction species with the active phase, leading to increased production of heavier hydrocarbons (higher C₅₊ selectivity).¹⁵⁵ On the contrary, the cobalt particles located outside the CNT channels exhibited lower activity and

C_{5+} selectivity in the FT reaction.¹⁵⁰ Comparison of the FT synthesis rates using graphene- and CNT-supported cobalt catalysts showed that the former increased the rate by 22%, shifted the product distribution to the long-chain hydrocarbons, and exhibited higher stability compared with the latter.¹⁵⁶ These properties were attributed to the better dispersion of cobalt clusters and an increase in the degree of reduction of Co at relatively lower temperatures in the graphene-supported catalyst.

Eschemann and co-workers demonstrated that by the choice of an appropriate solvent (alcohol) and drying procedure (under nitrogen flow) during the catalyst preparation step, clustering of supported nanoparticles could be avoided, yielding well-dispersed supported cobalt oxide nanoparticles.¹⁵⁷ It was found that the CNT-supported cobalt catalysts prepared by this method showed a superior cobalt-mass-based activity in FT synthesis compared with those prepared from an aqueous solution, as the cobalt particles had grown less during reduction and the catalytic test, leading to higher specific active metal surface area. CNT-supported cobalt catalysts exhibit superior activity and C_{5+} selectivity compared with their counterparts supported on $\gamma\text{-Al}_2\text{O}_3$ due to decreased Co particle size and increased cobalt reducibility, dispersion, and active metal surface area on CNTs.¹⁵⁸ Phaahlamohlaka et al.¹⁵⁹ utilized mesoporous hollow carbon spheres (MHCSs) as a model support to study the influence of Ru nanoparticle location relative to Co_3O_4 nanoparticles on the reduction behavior and activity of Co catalysts in FT synthesis. In this study, Ru nanoparticles were loaded both inside and outside the MHCSs, while Co_3O_4 particles were loaded on the outside of the MHCSs (Figure 10). Greater reducibility at low

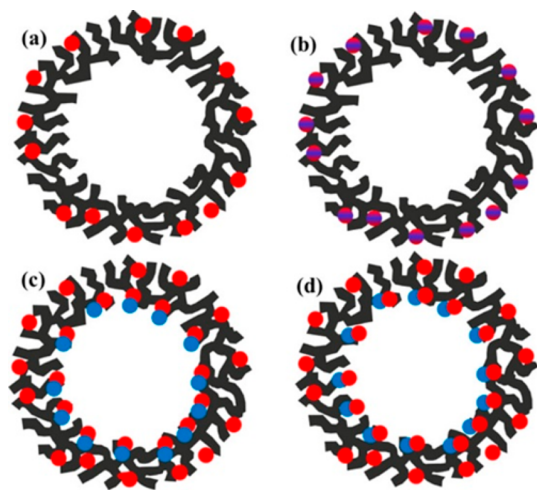


Figure 10. Schematic showing likely Co (red) and Ru (blue) particle distributions on the MHCS support: (a) Co/MHCS; (b) CoRu/MHCS; (c) 0.2% Ru@MHCS@Co; (d) 0.5% Ru@MHCS@Co.¹⁵⁹

temperatures and higher activity were observed for catalysts where the Co and Ru metals were in close proximity to each other (CoRu/MHCS, Figure 10b) as a result of a primary hydrogen spillover effect that enhanced the cobalt oxide reduction, leading to a highly active Co catalyst. In case of Ru@MHCS@Co, where the Co and Ru were separated by a carbon shell (Figure 10c,d), secondary hydrogen spillover induced complete reduction of cobalt oxide to metallic Co at a lower temperature compared with unpromoted catalyst.

In the direction of decreasing the multitude of parameters in the multistep FT catalytic preparation technique, plasma technology was implemented by Aluha and co-workers, enabling them to prepare C-supported nanometric Co catalysts in a single step.^{160–166} It was shown that C-supported nanometric Co catalysts synthesized by plasma spraying in a single step displayed superior catalytic activity (CO conversion: 20% vs 7%; process conditions: 40 wt % Co, 230 °C, 30 bar, $\text{H}_2/\text{CO} = 1.88$) over the conventional multistep catalyst production by precipitation or impregnation.¹⁶⁰ This approach to catalyst preparation resulted in nonporous catalysts and aided in improving the dispersion of cobalt compared with the impregnation or precipitation technique: the average Co cluster size in plasma-synthesized Co/C was ca. 12 nm, compared with 37 and 22 nm using the impregnation and precipitation techniques, respectively. The lower process temperature (220 °C) was more beneficial for the selectivity of the Co/C catalyst toward the gasoline ($C_4\text{--}C_{12}$) and diesel ($C_{13}\text{--}C_{20}$) range fractions, with CO conversion of about 42%. Increasing the process temperature to 260 °C resulted in complete CO conversion at the cost of producing highly prevalent undesired products such as CO_2 and CH_4 .^{161–163}

Extensive studies using carbon support materials derived directly from lignocellulosic biomass in FT synthesis are still missing. Such studies would help to identify suitable biomass resources from its diverse and heterogeneous pool of raw material. Furthermore, the presence of inorganic species in biomass-derived carbon supports might play a role in enhancing the activity in FT synthesis.

11. SILICON CARBIDE-SUPPORTED CATALYSTS

The reactions involved in the FT synthesis are highly exothermic and cause adiabatic temperature rises of up to 1477 °C (simply estimated as $\Delta H_r/C_{p,\text{reactants}}$).¹⁶⁷ Therefore, heat transfer and recovery is very important, especially when using a fixed-bed reactor. A local temperature rise in the reactor can form a hot spot in the catalytic bed that might lead to losses of activity and selectivity, sintering, and deactivation of the catalyst. For this reason, the Sasol I plant in South Africa used a small reactor tube with an internal diameter of 5 cm. The traditional FT catalytic supports such as alumina, silica, and titania have mediocre thermal conductivity and thus might not prevent the formation of hot spots within the catalyst bed.^{33,168} $\beta\text{-SiC}$ (SiC) is a covalent material made up of tetrahedral structural units. It is a newly emerging catalytic support material with very interesting properties such as chemical inertness, high thermal conductivity, and high mechanical strength along with surface area and surface nature that can be tuned by doping and open porosity consisting mostly of meso- and macropores.¹⁶⁹ Furthermore, it combines the best properties of both oxide- and carbon-based supports and circumvents many of their disadvantages.¹⁷⁰ The thermal conductivity of SiC gives rise to improved heat transfer in catalytic processes by reducing the formation of local hotspots as well as temperature gradients across the catalyst bed, making it a promising support material for FT catalysts. Furthermore, the chemical inertness of SiC leads to easy recovery of both the active phase and the support.¹⁷¹

Lacroix and co-workers evaluated the cobalt-based catalyst supported on a medium-surface-area SiC foam ceramic in FT synthesis (30 wt % Co, 220 °C, 40 bar, $\text{H}_2/\text{CO} = 2$) in a fixed-bed reactor with no catalyst diluent.¹⁴ The catalytic test results for these catalysts were compared with those obtained on a

reference Co/Al₂O₃ foam catalyst. It was found that at medium isoconversion (<50%) the two catalysts displayed similar C₅₊ selectivities. However, when the CO conversion was increased to 70%, a significant difference between the two catalysts in terms of the C₅₊ selectivity was observed (80% with Co/SiC vs 54% with Co/Al₂O₃). The selectivity difference was ascribed to the high efficiency of the thermally conductive SiC to evacuate heat generated during the course of the reaction and also to the presence of meso- and macroporosity in the SiC support. Furthermore, the hydrocarbon chain growth probability (α) obtained on the SiC-based catalyst (0.91) was higher than that of the counterpart on alumina (0.88) with preference to wax formation (Figure 11). Furthermore, the high chemical

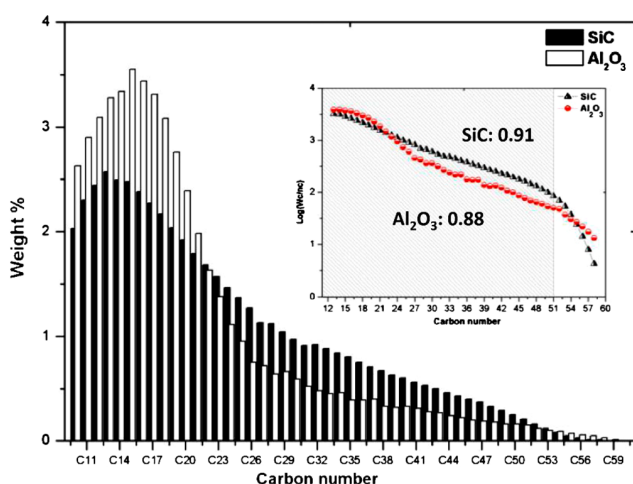


Figure 11. Distributions of liquid hydrocarbons obtained on the Co/SiC and Co/Al₂O₃ catalysts. The results were obtained from the liquid hydrocarbons trapped at high pressure in the hot trap (85 °C). Republished with permission from ref 14. Copyright 2011 Elsevier.

inertness of the SiC material allowed the recovery of both the active phase and the support by a simple acid washing. The recovered SiC impregnated with fresh Co catalyst displayed similar catalytic performance compared to the fresh catalyst. The catalytic activity of Co/SiC was further improved by supporting the cobalt phase on Al₂O₃-coated SiC (Co/Al₂O₃-SiC) without compromising its selectivity for C₅₊ hydrocarbons (80%). The improved activity was attributed to the higher cobalt dispersion on the Al₂O₃-SiC support compared with unmodified SiC.

In another study, Co/SiC catalysts promoted with Ru (to enhance Co reducibility) were evaluated in FT synthesis in a platelet-structured millireactor (PSR) at different gas hourly space velocities (GHSVs), and the catalytic results were compared with those obtained in the “conventional” micro-tubular fixed-bed reactor (TR).¹⁷² The PSR containing an open SiC cell foam morphology allowed easier operation under different GHSVs without problems of pressure drop or reactor plugging. The PSR displayed a better activity in FT synthesis (30 wt % Co, 220 °C, 10 bar, H₂/CO = 2) compared with that obtained in the TR, and the activity steadily increased with increasing GHSV. The improved catalytic activity with increasing GHSV in the PSR was ascribed to the efficient removal of the liquid hydrocarbon formed on the catalytic surface, ensuring better accessibility of the reactant to the active phase. Furthermore, the PSR yielded a narrower hydrocarbon distribution profile (between C₇ and C₁₇)

compared with the TR, which generated longer-chain hydrocarbons under similar reaction conditions. Of the cobalt precursors supported on SiC, namely, cobalt nitrate (CoN), cobalt acetate (CoA), cobalt chloride (CoCl), and Co citrate (CoCit), it was found that CoN/SiC showed the best activity in FT synthesis.¹⁷³ This was attributed to the larger Co particle size, the ease of Co reduction, and the basicity of CoN/SiC.

Diaz et al.¹⁷⁴ demonstrated that pretreatment of β -SiC with the pore agent and acidic treatment improved the catalytic activity of the resulting cobalt catalysts in the FT synthesis (13–15 wt % Co, 220 °C, 20 bar, H₂/CO = 2). The acid treatment resulted in SiC with lower metal impurities that were left behind during its preparation and also increased the number of acid sites. The pore agent treatment led to changes in the pore size distribution of SiC by increasing the macropore volume, which favored evacuation of the FT products, keeping the active sites available for reactants. Both these pretreatments enhanced the reducibility of cobalt particles with increased numbers of active sites, resulting in an improvement in the catalytic activity. Alternatively, Liu and co-workers evaluated cobalt catalysts supported on titania-decorated silicon carbide (β -SiC) in FT synthesis (10 wt % Co, 230 °C, 40 bar; H₂/CO = 2).¹⁷⁵ The catalytic test results showed that the cobalt time yield of the catalyst after TiO₂ modification ($7.5 \times 10^{-5} \text{ mol}_{\text{CO}} \text{ g}_{\text{Co}}^{-1} \text{ s}^{-1}$) was about 42% higher than that of the counterpart on the bare support ($5.3 \times 10^{-5} \text{ mol}_{\text{CO}} \text{ g}_{\text{Co}}^{-1} \text{ s}^{-1}$). However, the C₅₊ selectivities of the two catalysts were the same (ca. 91%). Two-dimensional (2D) elemental maps deduced by 2D energy-filtered TEM and ⁵⁹Co NMR analysis revealed that the nanoscale introduction of TiO₂ into the β -SiC matrix significantly enhanced the formation of small and medium-sized cobalt particles. Furthermore, the proper metal–support interaction between cobalt nanoparticles and TiO₂ led to the formation of smaller cobalt particles (<15 nm), resulting in a large fraction of surface atoms and, consequently, a great enhancement of the catalytic activity.

12. MISCELLANEOUS SUPPORTS

In addition to the supports discussed above, ceria, aluminum phosphate, hydrotalcites, and metal–organic frameworks (MOFs) as potential supports for cobalt-based catalysts have also been reported. In the following sections, the activities and selectivities of these catalysts in the FT reaction are discussed.

12.1. Aluminum Phosphate-Supported Catalysts. The surface of aluminum phosphate possesses acidic sites that are capable of participating in the catalytic process, and it could also be used as a catalytic carrier.¹⁷⁶ Amorphous AlPO₄ is built of tetrahedral AlO₄ and PO₄ units and is structurally similar to silica.¹⁷⁷ Furthermore, it exhibits a high surface area and large average pore size, which are beneficial characteristics of a support. Cobalt catalysts supported on AlPO₄ prepared from cobalt nitrate as the precursor displayed better catalytic performance but a slightly lower C₅₊ selectivity in FT synthesis compared with the corresponding Co/Al₂O₃ catalyst.¹⁷⁸ The enhanced activity was attributed to the improvement in the degree of Co reduction (at low temperature), with a high surface area of metallic cobalt containing proper electronic states that aid in the dissociative adsorption of CO, homogeneous dispersion of cobalt on AlPO₄, and decreased cobalt aluminate formation compared with the pristine Al₂O₃ support. The Co/AlPO₄ catalysts showed a rapid decrease in activity in the initial period, which was attributed to carbon

deposition, including filamentous carbons; however, the catalytic activity was still higher than that of Co/Al₂O₃.

12.2. Hydrotalcite-Supported Catalysts. Hydrotalcite-like compounds (HTIs), denoted by the empirical formula $[M^{2+}_{1-x}M^{3+}_x(OH)_2]^{x+}[A^{n-}_{x/n}]^{x-}\cdot mH_2O$, where M²⁺ is a divalent cation such as Co²⁺, Mg²⁺, Zn²⁺, Ni²⁺, or Cu²⁺; M³⁺ is a trivalent cation such as Al³⁺, Cr³⁺, Fe³⁺, or Ga³⁺; Aⁿ⁻ is an anion of charge n-; and m is the molar content of intercalated water, belong to the large class of anionic clays. In most cases, M²⁺ and M³⁺ are Mg²⁺ and Al³⁺, respectively. HTIs exhibit interesting properties for support materials such as high surface area, basic properties, and stability during thermal treatments.¹⁷⁹ Krylova et al.¹⁸⁰ demonstrated the use of synthetic MgAl-HTIs (Mg₆Al₂(CO₃)(OH)₁₆·4H₂O) as supports for Co (30 wt %) catalysts (Co/MgAl-HTIs) that are highly selective with respect to the formation of paraffins (α = 0.9–0.94) in FT synthesis (30 wt % Co, 170–230 °C, 10 bar, H₂/CO = 2). Another study found that Co (10 wt %) catalysts supported on commercial MgAl-HTIs displayed the highest catalytic activity and better C₅₊ selectivity than those supported on Al₂O₃ or MgO in FT synthesis (10 wt % Co, 230 °C, 2 bar, H₂/CO = 2).¹⁸¹ The better catalytic performance of Co/MgAl-HTIs was ascribed to the effects of the relatively higher total surface area, high reducibility, and smaller cobalt particle size. However, Co catalysts supported on calcined (500 °C) MgAl-HTIs displayed very poor activity, which was attributed to possible changes in the layered structure of HTIs, with the formation of Mg(Al) mixed oxide modifying the surface basicity and a much stronger interaction of Mg(Al) oxide with cobalt particles. Di Fronzo et al.¹⁸² introduced ternary HTIc, with the general formula $[Co_xZn_{(1-x-y)}Al_y(OH)_2](NO_3)_y\cdot 0.5H_2O$, synthesized by a modified urea method (Co–Zn–Al-HTIs with 5–35 wt % Co loading), as catalysts in FT synthesis. These are a new kind of catalyst, as the active metal is part of the structural core of the HTIc, as opposed to the conventional supported-cobalt catalysts. Cobalt was found to be uniformly dispersed in Co–Zn–Al-HTIs, and the reduction of cobalt was achieved at temperatures lower than 350 °C. These catalysts were tested for about 24 h and found to be active in FT synthesis (5 wt % Co, 220–240 °C, 20 bar, H₂/CO = 2), with CO conversions of about 80% and C₇₊ selectivities of up to ca. 93%. Although this is a new method of preparing Co-based FT catalysts, their long-term stability and the effect of cobalt loading in terms of morphology/size of Co crystallites must be established.

12.3. MOF-Supported Catalysts. MOFs are new class of porous organic–inorganic hybrid materials constructed by linking of metal ions or clusters with multidentate organic linkers. The tunable structure, topology, surface area, pore size, pore volume, and functionality of MOFs make them an interesting material for applications in the field of heterogeneous catalysis.^{183,184} Isaeva et al.¹⁸⁵ reported a cobalt (5–15 wt %)-containing Al³⁺-derived microporous MOF (Co-MIL-53(Al)) catalyst prepared by the incipient wetness impregnation method. Although MOFs are known for poor thermal stability (even <350 °C), these catalysts were found to retain their crystallinity even after the catalytic test (20 h). The Co-MIL-53(Al) catalysts exhibited similar activity and slightly higher C₅₊ selectivity (73% vs 69% at 15 wt % Co loading) than those supported on alumina in FT synthesis (5–15 wt % Co, 240 °C, 20 bar, H₂/CO = 2). Though Co nanocrystallites were more highly dispersed on MIL-53(Al) than on alumina, there was a significant fraction of small particles (2–4 nm) that

were not catalytically very active on the former support, which explains their similar catalytic performance. In a different study, MOF-derived cobalt catalysts, one containing nitrogen (Co-ZIF-67) and the other one nitrogen-free (Co-MOF-74), were prepared by incorporating cobalt into the MOF matrix during the synthesis.¹⁸⁶ Before these catalysts were tested in FT synthesis, they were thermally treated (550 °C) under an argon atmosphere to carbonize the organic ligands that aided in the reduction of cobalt species and to prevent Co aggregation. However, the MOF structure was collapsed after the thermal treatment. The cobalt nanocrystallites embedded in the carbon matrix were well-dispersed in both of these catalysts. The cobalt catalysts derived from nitrogen-free MOFs (Co-C) exhibited better CO conversion (30% vs 10%) with higher C₅₊ selectivity (65% vs 31%), moderate short-chain hydrocarbon selectivity (C₂–C₄, 10% vs 37%), and similar CH₄ selectivity compared with the catalysts derived from nitrogen-containing MOFs (Co-NC) in FT synthesis (30 wt % Co in Co–C and 52 wt % Co in Co-NC, 230 °C, 30 bar, H₂/CO = 2). The difference in the catalytic performances was attributed to the relatively large pore size (7 nm in Co–C vs 3 nm in Co-NC), which facilitates the diffusion of hydrocarbons. The high selectivity for short-chain hydrocarbons (C₂–C₄) using Co-NC was tentatively ascribed to the electronic effects generated by the presence of nitrogen with electron donor capability in the carbon matrix.

Usually, the cobalt phase present in MOF-derived materials is composed solely of fcc Co. However, the hcp Co phase is found to be more active in FT synthesis because CO molecules bind to and dissociate from the Co-hcp structure more readily than from the Co-fcc phase.^{118,187} Recently it was shown that well-dispersed Co-fcc nanoparticles (ca. 10 nm) were obtained on the porous-carbon-supported Co catalysts prepared by pyrolyzing a Co-MOF-71 precursor.¹⁸⁸ Furthermore, doping of Co-MOF-71 with variable quantities of the Si precursor Si(OC₂H₅)₄ and subsequent pyrolysis (250 °C) and hydrogenation (350 °C) treatments resulted in catalysts (Co–Si–C) with tunable Co-fcc nanoparticles that were partially transformed into Co-hcp phase via a Co₂C intermediate. The Co–Si–C catalysts holding high Co site density (mol_{active sites}/g_{cat}) showed very good activity and C₅₊ selectivity, with diesel fuels being the main constituents in the liquid product. The C₅₊ space-time yields (g_{hydrocarbons} g_{cat}⁻¹ h⁻¹) of the Co–Si–C catalysts were reported to be relatively higher than those of the conventional Co catalysts, but the tested conditions were different. Therefore, more studies to test the long-term stability and economic viability of preparing these types of catalysts are needed in order to comment on their application at large scale.

12.4. Metal-Foam-Supported Catalysts. Metal foams are a new kind of porous materials with a cellular structure that offer interesting characteristics for use as catalyst carriers, such as high thermal conductivity like that of SiC, limited detrimental metal–support interactions, a low pressure drop due to the presence of voids (macropores) that enhance the mass transfer properties, and low density.¹⁸⁹ Park et al.¹⁹⁰ introduced the use of Ni-foam-supported catalysts prepared by coating the cobalt catalyst powder (Co/Al₂O₃) on the surface of a Ni foam in FT synthesis (8 wt % Co, 190–220 °C, 20 bar, H₂/CO = 2). The thus-prepared Ni foam catalyst showed low CH₄ and CO₂ selectivities and high C₅₊ selectivity compared with the conventional Co/Al₂O₃ catalyst. The enhanced performance was attributed to the better heat and mass transfer properties in the reaction. Furthermore, it was found

Table 5. Guidance for Support Modifications To Improve Catalytic Performance of Co-Based Catalysts in FT Synthesis (⊗, No; ✓, Yes)

support	are surface modifications/doping required to improve the support characteristics?					
	surface area	reducibility	dispersion	metal–support interactions	acid/base properties	thermal conductivity
alumina	⊗	✓	✓	✓	✓	✓
silica	⊗	✓	✓	✓	✓	✓
titania	✓	⊗	✓	✓	✓	✓
niobia	✓	⊗	✓	✓	✓	✓
zirconia	✓	✓	✓	⊗	✓	✓
zeolites	⊗	✓	⊗	✓	⊗	✓
carbon materials	⊗	⊗	✓	⊗	✓	⊗
silicon carbide	✓	⊗	✓	⊗	✓	⊗
metal foams	✓	⊗	✓	⊗	✓	⊗

Table 6. Guidance for the Choice of Support Based on the Product Selectivities of Co-Based Catalysts in FT Synthesis

products	support	characteristics that affect the product selectivity	refs
C ₅₊	Al ₂ O ₃	<ul style="list-style-type: none"> • pore size (ca. 12–50 nm) • Co particle size (7–8 nm for γ-Al₂O₃) • from different phases of alumina, α-Al₂O₃ is preferred • catalyst particle size (50–225 μm) 	35, 36, 40
	surface-modified Al ₂ O ₃	<ul style="list-style-type: none"> • modification with inorganic oxides such as Zr, Ti, or Ta • alters H₂/CO mass transport within the pores • intermediate acid/base character ($\eta = 2.49$–2.53 eV; vide supra) 	49, 67
	SiO ₂	<ul style="list-style-type: none"> • pore size (ca. 10 nm) • Co particle size (ca. 9 nm) • moderate CO adsorption 	78, 79
	surface-modified SiO ₂	<ul style="list-style-type: none"> • modification with inorganic oxides such as Zr/P • enhances stabilization of the active phase • selective poisoning of methanation sites by carbon coating (10 wt %) 	81, 82, 89, 93
	TiO ₂	<ul style="list-style-type: none"> • reducible nature of the support • average pore size (ca. 16 nm) favors easy diffusion of long-chain hydrocarbons, leading to high C₅₊ selectivity • close interaction of Co promoted with Ru or Re • reduces blockage of the active site by carbonaceous species 	87, 107, 108
	Nb ₂ O ₅	<ul style="list-style-type: none"> • reducible nature of the support • electronic effect: possible electron enrichment of surface cobalt atoms by partially reduced NbO_x species localized on the cobalt particles favors hydrogenolysis, leading to higher C₅₊ selectivity 	113–115
	ZrO ₂	<ul style="list-style-type: none"> • ability to adsorb hydrogen via a spillover mechanism • mesoporous ZrO₂ with pore size of ca. 9 nm yields mostly C₁₂–C₁₈ (diesel range) 	10, 120
C ₅ –C ₁₀ (gasoline range)	carbon	<ul style="list-style-type: none"> • Co particle size (~8 nm) • CNTs: confinement of reactants and reaction intermediates within CNTs leads to increased contact time of reactive species with the active phase, favoring C₅₊ selectivity • graphite: reduces the high H adsorption and thus leads to long-chain growth 	146, 155, 156
	SiC	<ul style="list-style-type: none"> • ability to efficiently evacuate heat generated during the FT reaction, • presence of meso- and macroporosity favors wax formation even at high CO conversions • surface modification with TiO₂ leads to smaller Co particles (<15 nm) 	14, 175
	zeolites	<ul style="list-style-type: none"> • the acidic nature of the zeolites can be tuned by changing the silica/alumina ratio, which favors short-chain hydrocarbons • serve as bifunctional catalysts • secondary reaction of long-chain hydrocarbons to yield branched hydrocarbons with limited chain length and aromatics • preferred zeolites: ZSM-5 or hierarchical Y zeolite 	127–130, 137
oxygenates	CeO ₂	<ul style="list-style-type: none"> • redox properties of CeO₂ • alcohol formation based on Co–CeO₂ interfacial catalysis (vide supra) 	12, 141

that CH₄ formation was effectively minimized even at higher reaction temperatures. However, macroporous metal foams possess low metal surface area, making it difficult to get good dispersion of the active phase. In the direction of circumventing this problem, Harmel¹⁹¹ proposed a method of growing cobalt nanowires in the porous network of nickel and copper foams. Interestingly, these Co nanowires constitute Co-hcp phase, which is known to exhibit high activity in FT

synthesis. The Ni-foam-supported cobalt catalysts showed high activity and selectivity for CH₄, whereas the Cu-foam-supported catalysts exhibited high activity and selectivity for C₅₊ hydrocarbons. This represents a new method to increase the dispersion of the active phase and enhance the intrinsic activity of cobalt active sites, but the long-term stability of these catalysts and the economic viability of the preparation

protocol must still be carefully evaluated before they can be implemented in the industrial process.

12.5. Hydroxyapatite: A Potential Support Material.

Hydroxyapatite (HAP, $\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2$) is a double salt of tricalcium phosphate- and calcium hydroxide-based materials that contains both acidic and basic functionalities.¹⁹² Furthermore, it exhibits other characteristics such as high thermal stability, extremely low water solubility, and tunable specific surface area with or without mesopores depending on the synthesis conditions.¹⁹³ The Ca/P molar ratio in the stoichiometric form of HAP is 1.67. The ionic radii of HAP's component elements (Ca and P) permit the transfer or loss of ions within its crystal structure, consequently leading to nonstoichiometric HAPs, $\text{Ca}_{10-z}(\text{HPO}_4)_z(\text{PO}_4)_{6-z}(\text{OH})_{2-z}$ ($0 < z \leq 1$) with Ca/P molar ratios in the range of 1.50 to 1.80.¹⁹⁴ The possibility of preparing nonstoichiometric HAPs helps in tuning the density of acidic and basic sites, which is not possible in the case of conventional catalyst carriers. When the Ca/P ratio of HAP is about 1.50, the material behaves as an acidic catalyst. HAP exhibits characters of both acidic and basic catalysts when the Ca/P ratio is between 1.50 and 1.67 and acts as a basic catalyst when the Ca/P ratio is above 1.67.¹⁹⁵ Furthermore, the calcium ions in HAP can be easily exchanged with most divalent cations (e.g., Cu^{2+} , Pb^{2+} , Zn^{2+} , Cd^{2+} , Co^{2+} , etc.) without affecting the stability of the phosphate.¹⁹⁶ These interesting characteristics of HAPs make them unique support materials in heterogeneous catalysts.¹⁹⁷ Over the past decade, their potential as catalyst supports has been realized in many different reactions such as methane oxidation,^{198,199} dry reforming of methane,^{200–203} the water-gas shift reaction,²⁰⁴ hydrogen production,²⁰⁵ and selective oxidation of alcohols.²⁰⁶ Considering the properties of HAPs, it is definitely worth utilizing them as catalyst supports in Co-based FT catalysts.

13. SUMMARY

On the basis of the above-discussed literature on different aspects of support effects in the FT process, the guidelines for determining the necessity of support surface modifications are given in Table 5, and the choice of support on the basis of the product selectivities of Co-based catalysts in FT synthesis is summarized in Table 6.

14. CONCLUSIONS AND OUTLOOK

In general, cobalt precursors experience varying metal–support interactions depending on the type of the support. Oxide supports exhibit medium to strong interactions with cobalt, whereas pristine carbon materials and SiC display very poor interaction. The type of interaction between the support and the metal plays a pivotal role in the dispersion of cobalt metal precursors, and the strength and type of interaction determines the cobalt reducibility, which in turn determines the activity and selectivity in FT synthesis. Textural properties of the support such as surface area, pore volume, pore size distribution, and crystalline phase of the support have been found to influence both the activity and selectivity. High specific surface area and pore volume are required in order to increase the loading of active metal without compromising its dispersion. The pore size of the mesoporous supports limits the size of cobalt nanocrystallites via the confinement effect. Catalysts containing medium-sized pores exhibit higher selectivity for long-chain hydrocarbons compared with catalysts with small pores.

Al_2O_3 has been the most extensively studied support because of its applicability under industrial conditions. Different crystalline phases of alumina exhibit C_{5+} selectivities in the following order: $\alpha\text{-Al}_2\text{O}_3 > \delta\text{-Al}_2\text{O}_3 > \theta\text{-Al}_2\text{O}_3 > \gamma\text{-Al}_2\text{O}_3$. Different types of oxides (magnesia, zirconia, titania, phosphorus, lanthana), silanes, and organic molecules (sorbitol, cyclodextrin) have been used to modify the alumina surface to prevent the formation of refractory cobalt aluminates resulting from the direct interaction of cobalt with the alumina support. Porous $\gamma\text{-Al}_2\text{O}_3$ covered by monolayers of various oxides with identical support textures, extents of Co reduction, and particle sizes showed that the chemical nature of the support (Lewis acid/base character) controls both the catalytic activity and the selectivity for long-chain products.

Silica has been the next most widely studied support, and observations made for alumina in terms of textural properties hold true for silica supports too. Although the interaction of cobalt metal precursors with silica is poor compared with alumina, the formation of cobalt silicates is still observed. Such detrimental interactions are circumvented by surface modifications with inorganic oxides and organic compounds. Mesoporous silicas (MCM-41 and SBA-15) offer the possibility of synthesizing supports with desired narrow-pore size distributions, surface areas, pore volumes, and controllable acid/base properties.

Titania and niobia exhibit strong interactions with cobalt. These supports are partially reduced during the activation step under a hydrogen atmosphere and decorate the surface of the cobalt nanocrystallites. Zirconia shows weaker interactions with cobalt compared with the supports mentioned above and displays the capability of hydrogen adsorption via a spillover mechanism. Supports such as alumina, silica, and SiC can be surface-modified with titania, niobia, and zirconia to improve cobalt dispersion, reducibility, and stability, consequently enhancing the catalytic performance.

Zeolites with shape-selective characteristics, well-defined pore structures, high porosity, and acidity that can be tuned by changing the silica/alumina ratio. Cobalt catalysts supported on zeolites exhibit bifunctional characteristics that are very useful in directing the selectivity for the desired hydrocarbon fraction (e.g., gasoline vs diesel) through in situ hydrocracking of the wax fraction.

Carbon materials interact very weakly with cobalt, making them ideal supports for use in model catalysts for the rational design of catalysts. SiC support characteristics such as high thermal conductivity, mechanical strength, and chemical inertness make SiC a suitable support material for FT catalysts, as the FT reaction is highly exothermic. Advantages of using SiC-supported Co cobalt catalysts include improved heat transfer in FT synthesis, easy recovery of both the active phase and the support, and improved C_{5+} selectivity.

Significant advances have been realized in understanding the effect of the support in FT synthesis, but the majority of the studies have been performed on alumina and to a certain extent on silica. In order to mask unfavorable properties of the supports, surface modification by coating the support surface with inorganic oxides or organic molecules has been used to get desired properties that improve the activity and/or selectivity in FT synthesis. Parameters such as the cobalt nanocrystallite size, crystal structure (hcp vs fcc), distribution, and reducibility on a support material play a role in determining the activity of supported Co catalysts in FT synthesis. Some of the factors that influence the selectivities for

various products in FT synthesis are the pore structure and chemical nature of the support, the type of support material (selectivities for C₅₊ hydrocarbons: SiC > titania/ α -alumina > γ -alumina/silica), and the acidic character of the support, which induces in situ hydrocracking of long-chain hydrocarbons (e.g., in the case of zeolites). The FT catalysts prepared on unconventional supports such as mesoporous silicas, zeolites, SiC, MOFs, hydrotalcite, and metal foams present enhanced activity and/or selectivity for desired hydrocarbon ranges in FT synthesis. However, to implement these new findings in industry, more work is needed to establish the long-term operational stability, recovery and regenerability, cost effectiveness, and environmental integrity of these catalysts. In particular, through the use of advanced electron microscopy techniques such as cryo-TEM and electron tomography, more detailed information on the distribution of the metal precursors can be obtained at different stages of catalyst preparation with the newly emerging supports. Furthermore, in situ X-ray absorption spectroscopy, X-ray fluorescence spectroscopy, and liquid-cell TEM studies would provide a quantitative picture of the evolution of the diameter, shape, lattice packing, and density of the deposited cobalt particles in real time. In order to design supported catalysts with the desired activity and selectivity, emerging in situ spectroscopy and microscopy techniques under industrial working conditions would help in better understanding the role of electronic and structural effects in the FT reaction. Finally, the development of next-generation supported cobalt catalysts that resist some impurities in syngas derived from biomass and wastes must be intensified. New potential supports such as hydroxyapatite may bring new insights for the design of competitive catalysts versus traditional alumina- or silica-based supports in FT synthesis.

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Notes

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REFERENCES

- Anderson, R. B.; Shultz, J. F.; Hofer, L. J. E.; Storch, H. H. *Physical Chemistry of the Fischer–Tropsch Synthesis*; Bureau of Mines, Bulletin 580; U.S. Department of the Interior: Washington, DC, 1959.
- Iglesia, E. Design, Synthesis, and Use of Cobalt-Based Fischer–Tropsch Synthesis Catalysts. *Appl. Catal., A* **1997**, *161* (1–2), 59–78.
- Dry, M. E. Fischer–Tropsch Synthesis over Iron Catalysts. *Catal. Lett.* **1990**, *7* (1–4), 241–251.
- Davis, B. H. Fischer–Tropsch Synthesis: Comparison of Performances of Iron and Cobalt Catalysts. *Ind. Eng. Chem. Res.* **2007**, *46* (26), 8938–8945.
- Khodakov, A. Y.; Chu, W.; Fongarland, P. Advances in the Development of Novel Cobalt Fischer–Tropsch Catalysts for Synthesis of Long-Chain Hydrocarbons and Clean Fuels. *Chem. Rev.* **2007**, *107*, 1692–1744.
- van de Loosdrecht, J.; Botes, F. G.; Ciobica, I. M.; Ferreira, A.; Gibson, P.; Moodley, D. J.; Saib, A. M.; Visagie, J. L.; Weststrate, C. J.; Niemantsverdriet, J. W. Fischer–Tropsch Synthesis: Catalysts and Chemistry. In *Comprehensive Inorganic Chemistry II*, 2nd ed.; Reedijk, J., Poepelmeier, K., Eds.; Elsevier: Amsterdam, 2013; Vol. 7, Chapter 7.20.
- Oukaci, R.; Singleton, A. H.; Goodwin, J. G. Comparison of Patented Co F–T Catalysts Using Fixed-Bed and Slurry Bubble Column Reactors. *Appl. Catal., A* **1999**, *186* (1–2), 129–144.
- Sun, B.; Xu, K.; Nguyen, L.; Qiao, M.; Tao, F. F. Preparation and Catalysis of Carbon-Supported Iron Catalysts for Fischer–Tropsch Synthesis. *ChemCatChem* **2012**, *4*, 1498–1511.
- Khodakov, A. Y.; Zholobenko, V. L.; Bechara, R.; Durand, D. Impact of Aqueous Impregnation on the Long-Range Ordering and Mesoporous Structure of Cobalt Containing MCM-41 and SBA-15 Materials. *Microporous Mesoporous Mater.* **2005**, *79* (1–3), 29–39.
- Enache, D. I.; Roy-Auberger, M.; Revel, R. Differences in the Characteristics and Catalytic Properties of Cobalt-Based Fischer–Tropsch Catalysts Supported on Zirconia and Alumina. *Appl. Catal., A* **2004**, *268* (1–2), 51–60.
- den Otter, J. H.; de Jong, K. P. Highly Selective and Active Niobia-Supported Cobalt Catalysts for Fischer–Tropsch Synthesis. *Top. Catal.* **2014**, *57* (6–9), 445–450.
- Spadaro, L.; Arena, F.; Granados, M. L.; Ojeda, M.; Fierro, J. L. G.; Frusteri, F. Metal-Support Interactions and Reactivity of Co/CeO₂ Catalysts in the Fischer–Tropsch Synthesis Reaction. *J. Catal.* **2005**, *234* (2), 451–462.
- Bessell, S. Investigation of Bifunctional Zeolite Supported Cobalt Fischer–Tropsch Catalysts. *Appl. Catal., A* **1995**, *126* (2), 235–244.
- Lacroix, M.; Dreibine, L.; De Tymowski, B.; Vigneron, F.; Edouard, D.; Bégin, D.; Nguyen, P.; Pham, C.; Savin-Poncet, S.; Luck, F.; Ledoux, M. J.; Pham-Huu, C. Silicon Carbide Foam Composite Containing Cobalt as a Highly Selective and Re-Usable Fischer–Tropsch Synthesis Catalyst. *Appl. Catal., A* **2011**, *397* (1–2), 62–72.
- Diehl, F.; Khodakov, A. Y. Promotion of Cobalt Fischer–Tropsch Catalysts with Noble Metals: A Review. *Oil Gas Sci. Technol.* **2009**, *64* (1), 11–24.
- Morales, F.; Weckhuysen, B. M. Promotion Effects in Co-Based Fischer–Tropsch Catalysis. *Catalysis* **2006**, *19*, 1–40.
- Yang, N.; Bent, S. F. Investigation of Inherent Differences between Oxide Supports in Heterogeneous Catalysis in the Absence of Structural Variations. *J. Catal.* **2017**, *351*, 49–58.
- Reuel, R. C.; Bartholomew, C. H. The Stoichiometries of H₂ and CO Adsorptions on Cobalt: Effects of Support and Preparation. *J. Catal.* **1984**, *85* (1), 63–77.
- Jahangiri, H.; Bennett, J.; Mahjoubi, P.; Wilson, K.; Gu, S. A Review of Advanced Catalyst Development for Fischer–Tropsch Synthesis of Hydrocarbons from Biomass Derived Syn-Gas. *Catal. Sci. Technol.* **2014**, *4* (8), 2210–2229.
- Zowtiak, J. M.; Bartholomew, C. H. The Kinetics of H₂ Adsorption on and Desorption from Cobalt and the Effects of Support Thereon. *J. Catal.* **1983**, *83* (1), 107–120.
- Bessell, S. Support Effects in Cobalt-Based Fischer–Tropsch Catalysis. *Appl. Catal., A* **1993**, *96* (2), 253–268.
- Bartholomew, C. H.; Reuel, R. C. Cobalt–Support Interactions: Their Effects on Adsorption and CO Hydrogenation Activity and Selectivity Properties. *Ind. Eng. Chem. Prod. Res. Dev.* **1985**, *24* (1), 56–61.
- Bertole, C. J.; Mims, C. A.; Kiss, G. Support and Rhenium Effects on the Intrinsic Site Activity and Methane Selectivity of Cobalt Fischer–Tropsch Catalysts. *J. Catal.* **2004**, *221* (1), 191–203.
- Zhang, Q.; Kang, J.; Wang, Y. Development of Novel Catalysts for Fischer–Tropsch Synthesis: Tuning the Product Selectivity. *ChemCatChem* **2010**, *2*, 1030–1058.
- Rytter, E.; Tsakoumis, N. E.; Holmen, A. On the Selectivity to Higher Hydrocarbons in Co-Based Fischer–Tropsch Synthesis. *Catal. Today* **2016**, *261*, 3–16.
- Rytter, E.; Holmen, A. On the Support in Cobalt Fischer–Tropsch Synthesis—Emphasis on Alumina and Aluminates. *Catal. Today* **2016**, *275*, 11–19.
- Ail, S. S.; Dasappa, S. Biomass to Liquid Transportation Fuel via Fischer–Tropsch Synthesis - Technology Review and Current Scenario. *Renewable Sustainable Energy Rev.* **2016**, *58*, 267–286.

- (28) Althapp, A.; Degen, W.; Heinel, E.; Rouveiroles, P.; Rudloff, M.; Stöhr, M. Renewable Fuels for Advanced Powertrains (RENEW), EC Contract No. SES6-CT-2003-502705, 2007.
- (29) Busca, G. The Surface of Transitional Aluminas: A Critical Review. *Catal. Today* **2014**, *226*, 2–13.
- (30) Wang, W. J.; Chen, Y. W. Influence of Metal Loading on the Reducibility and Hydrogenation Activity of Cobalt/Alumina Catalysts. *Appl. Catal.* **1991**, *77* (2), 223–233.
- (31) Chin, R. L.; Hercules, D. M. Surface Spectroscopic Characterization of Cobalt-Alumina Catalysts. *J. Phys. Chem.* **1982**, *86* (3), 360–367.
- (32) Jacobs, G.; Das, T. K.; Zhang, Y.; Li, J.; Racoillet, G.; Davis, B. H. Fischer–Tropsch Synthesis: Support, Loading, and Promoter Effects on the Reducibility of Cobalt Catalysts. *Appl. Catal., A* **2002**, *233* (1–2), 263–281.
- (33) Storsæter, S.; Tøtdal, B.; Walmsley, J. C.; Tanem, B. S.; Holmen, A. Characterization of Alumina-, Silica-, and Titania-Supported Cobalt Fischer–Tropsch Catalysts. *J. Catal.* **2005**, *236* (1), 139–152.
- (34) Xiong, H.; Zhang, Y.; Wang, S.; Li, J. Fischer–Tropsch Synthesis: The Effect of Al₂O₃ Porosity on the Performance of Co/Al₂O₃ Catalyst. *Catal. Commun.* **2005**, *6* (8), 512–516.
- (35) Borg, Ø.; Eri, S.; Blekkan, E. A.; Storsæter, S.; Wigum, H.; Rytter, E.; Holmen, A. Fischer–Tropsch Synthesis over γ -Alumina-Supported Cobalt Catalysts: Effect of Support Variables. *J. Catal.* **2007**, *248* (1), 89–100.
- (36) Borg, Ø.; Dietzel, P. D. C.; Spjelkavik, A. I.; Tveten, E. Z.; Walmsley, J. C.; Diplas, S.; Eri, S.; Holmen, A.; Rytter, E. Fischer–Tropsch Synthesis: Cobalt Particle Size and Support Effects on Intrinsic Activity and Product Distribution. *J. Catal.* **2008**, *259* (2), 161–164.
- (37) Schanke, D.; Eri, S.; Rytter, E.; Aaserud, C.; Hilmen, A.-M.; Lindvåg, O. A.; Bergene, E.; Holmen, A. Fischer–Tropsch Synthesis on Cobalt Catalysts Supported on Different Aluminas. *Stud. Surf. Sci. Catal.* **2004**, *147*, 301–306.
- (38) Phan, X. K.; Yang, J.; Bakhtiyar-Davijnay, H.; Myrstad, R.; Venvik, H. J.; Holmen, A. Studies of Macroporous Structured Alumina Based Cobalt Catalysts for Fischer–Tropsch Synthesis. *Catal. Lett.* **2011**, *141* (12), 1739–1745.
- (39) Vosoughi, V.; Badoga, S.; Dalai, A. K.; Abatzoglou, N. Modification of Mesoporous Alumina as a Support for Cobalt-Based Catalyst in Fischer–Tropsch Synthesis. *Fuel Process. Technol.* **2017**, *162*, 55–65.
- (40) Rane, S.; Borg, Ø.; Yang, J.; Rytter, E.; Holmen, A. Effect of Alumina Phases on Hydrocarbon Selectivity in Fischer–Tropsch Synthesis. *Appl. Catal., A* **2010**, *388* (1–2), 160–167.
- (41) Liu, Y.; Jia, L.; Hou, B.; Sun, D.; Li, D. Cobalt Aluminate-Modified Alumina as a Carrier for Cobalt in Fischer–Tropsch Synthesis. *Appl. Catal., A* **2017**, *530*, 30–36.
- (42) Zhang, Y.; Xiong, H.; Liew, K.; Li, J. Effect of Magnesia on Alumina-Supported Cobalt Fischer–Tropsch Synthesis Catalysts. *J. Mol. Catal. A: Chem.* **2005**, *237* (1–2), 172–181.
- (43) Chernavskii, P. A.; Pankina, G. A.; Lunin, V. V. The Influence of Oxide-Oxide Interaction on the Catalytic Properties of Co/Al₂O₃ in CO Hydrogenation. *Catal. Lett.* **2000**, *66* (3), 121–124.
- (44) Enger, B. C.; Fossan, Å. L.; Borg, Ø.; Rytter, E.; Holmen, A. Modified Alumina as Catalyst Support for Cobalt in the Fischer–Tropsch Synthesis. *J. Catal.* **2011**, *284* (1), 9–22.
- (45) Balonek, C. M.; Lillebø, A. H.; Rane, S.; Rytter, E.; Schmidt, L. D.; Holmen, A. Effect of Alkali Metal Impurities on Co-Re Catalysts for Fischer–Tropsch Synthesis from Biomass-Derived Syngas. *Catal. Lett.* **2010**, *138* (1–2), 8–13.
- (46) Bao, A.; Li, J.; Zhang, Y. Effect of Barium on Reducibility and Activity for Cobalt-Based Fischer–Tropsch Synthesis Catalysts. *J. Nat. Gas Chem.* **2010**, *19* (6), 622–627.
- (47) Xiong, H.; Zhang, Y.; Liew, K.; Li, J. Catalytic Performance of Zirconium-Modified Co/Al₂O₃ for Fischer–Tropsch Synthesis. *J. Mol. Catal. A: Chem.* **2005**, *231* (1–2), 145–151.
- (48) Rohr, F.; Lindvåg, O. A.; Holmen, A.; Blekkan, E. A. Fischer–Tropsch Synthesis over Cobalt Catalysts Supported on Zirconia-Modified Alumina. *Catal. Today* **2000**, *58* (4), 247–254.
- (49) Ma, W.; Jacobs, G.; Gao, P.; Jermwongratanchai, T.; Shafer, W. D.; Pendyala, V. R. R.; Yen, C. H.; Klettlinger, J. L. S.; Davis, B. H. Fischer–Tropsch Synthesis: Pore Size and Zr Promotional Effects on the Activity and Selectivity of 25%Co/Al₂O₃ Catalysts. *Appl. Catal., A* **2014**, *475*, 314–324.
- (50) Park, S. J.; Kim, S. M.; Woo, M. H.; Bae, J. W.; Jun, K. W.; Ha, K. S. Effects of Titanium Impurity on Alumina Surface for the Activity of Co/Ti-Al₂O₃ Fischer–Tropsch Catalyst. *Appl. Catal., A* **2012**, *419–420* (420), 148–155.
- (51) Liu, Y.; Luo, J.; Girleanu, M.; Ersen, O.; Pham-Huu, C.; Meny, C. Efficient Hierarchically Structured Composites Containing Cobalt Catalyst for Clean Synthetic Fuel Production from Fischer–Tropsch Synthesis. *J. Catal.* **2014**, *318*, 179–192.
- (52) Ledford, J. S.; Houalla, M.; Petrakis, L.; Hercules, D. M. Influence of Lanthanum Oxide on the Surface Structure and Co Hydrogenation Activity of Supported Cobalt Catalysts. *Stud. Surf. Sci. Catal.* **1987**, *31* (C), 433–442.
- (53) Cai, Z.; Li, J.; Liew, K.; Hu, J. Effect of La₂O₃-Doping on the Al₂O₃ Supported Cobalt Catalyst for Fischer–Tropsch Synthesis. *J. Mol. Catal. A: Chem.* **2010**, *330* (1–2), 10–17.
- (54) Brabant, C.; Khodakov, A.; Griboval-Constant, A. Promotion of Lanthanum-Supported Cobalt-Based Catalysts for the Fischer–Tropsch Reaction. *C. R. Chim.* **2017**, *20* (1), 40–46.
- (55) Iwamoto, R.; Grimblot, J. Influence of Phosphorus on the Properties of Alumina-Based Hydrotreating Catalysts. *Adv. Catal.* **1999**, *44* (C), 417–503.
- (56) Bae, J. W.; Kim, S. M.; Lee, Y. J.; Lee, M. J.; Jun, K. W. Enhanced Fischer–Tropsch Activity on Co/P-Al₂O₃ Catalyst: Effect of Phosphorous Content. *Catal. Commun.* **2009**, *10* (9), 1358–1362.
- (57) Woo, M. H.; Cho, J. M.; Jun, K. W.; Lee, Y. J.; Bae, J. W. Thermally Stabilized Cobalt-Based Fischer–Tropsch Catalysts by Phosphorous Modification of Al₂O₃: Effect of Calcination Temperatures on Catalyst Stability. *ChemCatChem* **2015**, *7* (9), 1460–1469.
- (58) Park, S. J.; Bae, J. W.; Jung, G. I.; Ha, K. S.; Jun, K. W.; Lee, Y. J.; Park, H. G. Crucial Factors for Catalyst Aggregation and Deactivation on Co/Al₂O₃ in a Slurry-Phase Fischer–Tropsch Synthesis. *Appl. Catal., A* **2012**, *413–414*, 310–321.
- (59) Ha, K. S.; Jung, G. I.; Woo, M. H.; Jun, K. W.; Bae, J. W. Effects of Phosphorus and Saccharide on Size, Shape, and Reducibility of Fischer–Tropsch Catalysts for Slurry Phase and Fixed-Bed Reactions. *Appl. Catal., A* **2013**, *453*, 358–369.
- (60) Park, S. J.; Bae, J. W.; Lee, Y. J.; Ha, K. S.; Jun, K. W.; Karandikar, P. Deactivation Behaviors of Pt or Ru Promoted Co/P-Al₂O₃ Catalysts during Slurry-Phase Fischer–Tropsch Synthesis. *Catal. Commun.* **2011**, *12* (6), 539–543.
- (61) Vissers, J. P. R.; Mercx, F. P. M.; Bouwens, S. M. A. M.; de Beer, V. H. J.; Prins, R. Carbon-Covered Alumina as a Support for Sulfide Catalysts. *J. Catal.* **1988**, *114* (2), 291–302.
- (62) Jean-Marie, A.; Griboval-Constant, A.; Khodakov, A. Y.; Diehl, F. Influence of Sub-Stoichiometric Sorbitol Addition Modes on the Structure and Catalytic Performance of Alumina-Supported Cobalt Fischer–Tropsch Catalysts. *Catal. Today* **2011**, *171* (1), 180–185.
- (63) Jean-Marie, A.; Griboval-Constant, A.; Khodakov, A. Y.; Monflier, E.; Diehl, F. β -Cyclodextrin for Design of Alumina Supported Cobalt Catalysts Efficient in Fischer–Tropsch Synthesis. *Chem. Commun.* **2011**, *47* (38), 10767–10769.
- (64) Rytter, E.; Salman, A. u. R.; Tsakoumis, N. E.; Myrstad, R.; Yang, J.; Lögberg, S.; Holmen, A.; Rønning, M. Hydrophobic Catalyst Support Surfaces by Silylation of γ -Alumina for Co/Re Fischer–Tropsch Synthesis. *Catal. Today* **2018**, *299*, 20–27.
- (65) Chen, L.; Tian, X.; Fu, Y.; Shen, J. The Effect of Surface Acidic and Basic Properties of Highly Loaded Co Catalysts on the Fischer–Tropsch Synthesis. *Catal. Commun.* **2012**, *28* (3), 155–158.
- (66) Zhang, J.; Chen, J.; Ren, J.; Li, Y.; Sun, Y. Support Effect of Co/Al₂O₃ Catalysts for Fischer–Tropsch Synthesis. *Fuel* **2003**, *82* (5), 581–586.

- (67) Prieto, G.; De Mello, M. I. S.; Concepción, P.; Murciano, R.; Pergher, S. B. C.; Martínez, A. Cobalt-Catalyzed Fischer–Tropsch Synthesis: Chemical Nature of the Oxide Support as a Performance Descriptor. *ACS Catal.* **2015**, *5* (6), 3323–3335.
- (68) Jeong, N. C.; Lee, J. S.; Tae, E. L.; Lee, Y. J.; Yoon, K. B. Acidity Scale for Metal Oxides and Sanderson's Electronegativities of Lanthanide Elements. *Angew. Chem., Int. Ed.* **2008**, *47* (52), 10128–10132.
- (69) Soled, S. L.; Iglesia, E.; Fiato, R. A.; Baumgartner, J. E.; Vroman, H.; Miseo, S. Control of Metal Dispersion and Structure by Changes in the Solid-State Chemistry of Supported Cobalt Fischer–Tropsch Catalysts. *Top. Catal.* **2003**, *26* (1–4), 101–109.
- (70) Borg, Ø.; Rønning, M.; Storster, S.; van Beek, W.; Holmen, A. Identification of Cobalt Species during Temperature Programmed Reduction of Fischer–Tropsch Catalysts. *Stud. Surf. Sci. Catal.* **2007**, *163*, 255–272.
- (71) Puskas, I.; Fleisch, T. H.; Hall, J. B.; Meyers, B. L.; Roginski, R. T. Metal-Support Interactions in Precipitated, Magnesium-Promoted Cobalt-Silica Catalysts. *J. Catal.* **1992**, *134* (2), 615–628.
- (72) Rosynek, M. P.; Polansky, C. A. Effect of Cobalt Source on the Reduction Properties of Silica-Supported Cobalt Catalysts. *Appl. Catal.* **1991**, *73* (1), 97–112.
- (73) Huber, G. W.; Guymon, C. G.; Conrad, T. L.; Stephenson, B. C.; Bartholomew, C. H. Hydrothermal Stability of Co/SiO₂ Fischer–Tropsch Synthesis Catalysts. *Stud. Surf. Sci. Catal.* **2001**, *139*, 423–430.
- (74) Ernst, B.; Libs, S.; Chaumette, P.; Kiennemann, A. Preparation and Characterization of Fischer–Tropsch Active Co/SiO₂ Catalysts. *Appl. Catal., A* **1999**, *186* (1–2), 145–168.
- (75) Khodakov, A. Y.; Lynch, J.; Bazin, D.; Rebours, B.; Zanier, N.; Moisson, B.; Chaumette, P. Reducibility of Cobalt Species in Silica-Supported Fischer–Tropsch Catalysts. *J. Catal.* **1997**, *168* (1), 16–25.
- (76) Tsubaki, N.; Sun, S.; Fujimoto, K. Different Functions of the Noble Metals Added to Cobalt Catalysts for Fischer–Tropsch Synthesis. *J. Catal.* **2001**, *199* (2), 236–246.
- (77) Lapszewicz, J. A.; Loeh, H. J.; Chipperfield, J. R. The Effect of Catalyst Porosity on Methane Selectivity in the Fischer–Tropsch Reaction. *J. Chem. Soc., Chem. Commun.* **1993**, No. 11, 913.
- (78) Saib, A. M.; Claeys, M.; Van Steen, E. Silica Supported Cobalt Fischer–Tropsch Catalysts: Effect of Pore Diameter of Support. *Catal. Today* **2002**, *71* (3–4), 395–402.
- (79) Song, D.; Li, J. Effect of Catalyst Pore Size on the Catalytic Performance of Silica Supported Cobalt Fischer–Tropsch Catalysts. *J. Mol. Catal. A: Chem.* **2006**, *247* (1–2), 206–212.
- (80) All, S.; Chen, B.; Goodwin, J. G. Zr Promotion of Co/SiO₂ for Fischer–Tropsch Synthesis. *J. Catal.* **1995**, *157* (1), 35–41.
- (81) Feller, A.; Claeys, M.; van Steen, E. Cobalt Cluster Effects in Zirconium Promoted Co/SiO₂ Fischer–Tropsch Catalysts. *J. Catal.* **1999**, *185* (1), 120–130.
- (82) Moradi, G. R.; Basir, M. M.; Taeb, A.; Kiennemann, A. Promotion of Co/SiO₂ Fischer–Tropsch Catalysts with Zirconium. *Catal. Commun.* **2003**, *4* (1), 27–32.
- (83) Bae, J. W.; Park, S. J.; Woo, M. H.; Cheon, J. Y.; Ha, K. S.; Jun, K. W.; Lee, D. H.; Jung, H. M. Enhanced Catalytic Performance by Zirconium Phosphate-Modified SiO₂-Supported Ru, Co Catalyst for Fischer–Tropsch Synthesis. *ChemCatChem* **2011**, *3* (8), 1342–1347.
- (84) Jang, I. H.; Um, S. H.; Lim, B.; Woo, M. H.; Jun, K. W.; Lee, J. B.; Bae, J. W. Effects of Surface Modification with Zirconium Phosphate on Ru/Co/SiO₂ Fischer–Tropsch Catalysts Analyzed by XPS and TEM Analyses. *Appl. Catal., A* **2013**, *450*, 88–95.
- (85) Jiang, Y.; Fu, T.; Lü, J.; Li, Z. A Zirconium Modified Co/SiO₂ Fischer–Tropsch Catalyst Prepared by Dielectric-Barrier Discharge Plasma. *J. Energy Chem.* **2013**, *22* (3), 506–511.
- (86) Hinchiranan, S.; Zhang, Y.; Nagamori, S.; Vitidsant, T.; Tsubaki, N. TiO₂ Promoted Co/SiO₂ Catalysts for Fischer–Tropsch Synthesis. *Fuel Process. Technol.* **2008**, *89* (4), 455–459.
- (87) Iglesia, E.; Soled, S. L.; Fiato, R. A.; Via, G. H. Bimetallic Synergy in Cobalt–Ruthenium Fischer–Tropsch Synthesis Catalysts. *J. Catal.* **1993**, *143* (2), 345–368.
- (88) den Otter, J. H.; Nijveld, S. R.; de Jong, K. P. Synergistic Promotion of Co/SiO₂ Fischer–Tropsch Catalysts by Niobia and Platinum. *ACS Catal.* **2016**, *6* (3), 1616–1623.
- (89) Gnanamani, M. K.; Jacobs, G.; Graham, U. M.; Pendyala, V. R. R.; Martinelli, M.; MacLennan, A.; Hu, Y.; Davis, B. H. Effect of Sequence of P and Co Addition over Silica for Fischer–Tropsch Synthesis. *Appl. Catal., A* **2017**, *538*, 190–198.
- (90) Gnanamani, M. K.; Jacobs, G.; Pendyala, V. R. R.; Graham, U. M.; Hopps, S. D.; Thomas, G. A.; Shafer, W. D.; Sparks, D. E.; Xiao, Q.; Hu, Y.; Davis, B. H. Fischer–Tropsch Synthesis: Anchoring of Cobalt Particles in Phosphorus Modified Cobalt/Silica Catalysts. *Appl. Catal., A* **2016**, *523*, 146–158.
- (91) Zhang, Y.; Hanayama, K.; Tsubaki, N. The Surface Modification Effects of Silica Supported by Organic Solvents for Fischer–Tropsch Synthesis Catalysts. *Catal. Commun.* **2006**, *7* (5), 251–254.
- (92) Chen, J.; Zhang, Y.; Tan, L.; Zhang, Y. A Simple Method for Preparing the Highly Dispersed Supported Co₃O₄ on Silica Support. *Ind. Eng. Chem. Res.* **2011**, *50* (7), 4212–4215.
- (93) Cheng, K.; Subramanian, V.; Carvalho, A.; Ordonsky, V. V.; Wang, Y.; Khodakov, A. Y. The Role of Carbon Pre-Coating for the Synthesis of Highly Efficient Cobalt Catalysts for Fischer–Tropsch Synthesis. *J. Catal.* **2016**, *337*, 260–271.
- (94) Allothman, Z. A. A Review: Fundamental Aspects of Silicate Mesoporous Materials. *Materials* **2012**, *5* (12), 2874–2902.
- (95) Beck, J. S.; Vartuli, J. C.; Roth, W. J.; Leonowicz, M. E.; Kresge, C. T.; Schmitt, K. D.; Chu, C. T.-W.; Olson, D. H.; Sheppard, E. W.; McCullen, S. B.; Higgins, J. B.; Schlenker, J. L. A New Family of Mesoporous Molecular Sieves Prepared with Liquid Crystal Template. *J. Am. Chem. Soc.* **1992**, *114* (27), 10834–10843.
- (96) Zhao, D.; Feng, J.; Huo, Q.; Melosh, N.; Fredrickson, G. H.; Chmelka, B. F.; Stucky, G. D. Triblock Copolymer Syntheses of Mesoporous Silica with Periodic 50 to 300 Angstrom Pores. *Science* **1998**, *279* (5350), 548–552.
- (97) Khodakov, A. Y.; Griboval-Constant, A.; Bechara, R.; Zholobenko, V. L. Pore Size Effects in Fischer–Tropsch Synthesis over Cobalt-Supported Mesoporous Silicas. *J. Catal.* **2002**, *206* (2), 230–241.
- (98) González, O.; Pérez, H.; Navarro, P.; Almeida, L. C.; Pacheco, J. G.; Montes, M. Use of Different Mesostructured Materials Based on Silica as Cobalt Supports for the Fischer–Tropsch Synthesis. *Catal. Today* **2009**, *148* (1–2), 140–147.
- (99) Khodakov, A. Y.; Bechara, R.; Griboval-Constant, A. Fischer–Tropsch Synthesis over Silica Supported Cobalt Catalysts: Mesoporous Structure versus Cobalt Surface Density. *Appl. Catal., A* **2003**, *254* (2), 273–288.
- (100) Tauster, S. J.; Fung, S. C.; Garten, R. L. Strong Metal–Support Interactions. Group 8 Noble Metals Supported on TiO₂. *J. Am. Chem. Soc.* **1978**, *100* (1), 170–175.
- (101) Resasco, D. E.; Weber, R. S.; Sakellson, S.; McMillan, M.; Haller, G. L. X-Ray Absorption near-Edge Structure Evidence for Direct Metal–Metal Bonding and Electron Transfer in Reduced Rhodium/Titanium Catalysts. *J. Phys. Chem.* **1988**, *92* (1), 189–193.
- (102) Beard, B. C.; Ross, P. N. Platinum–Titanium Alloy Formation from High-Temperature Reduction of a Titania-Impregnated Platinum Catalyst: Implications for Strong Metal–Support Interaction. *J. Phys. Chem.* **1986**, *90* (26), 6811–6817.
- (103) de la Peña O'Shea, V. A.; Consuelo Álvarez Galván, M.; Platero Prats, A. E.; Campos-Martin, J. M.; Fierro, J. L. G. Direct Evidence of the SMSI Decoration Effect: The Case of Co/TiO₂ Catalyst. *Chem. Commun.* **2011**, *47* (25), 7131.
- (104) Cats, K. H.; Andrews, J. C.; Stéphan, O.; March, K.; Karunakaran, C.; Meirer, F.; de Groot, F. M. F.; Weckhuysen, B. M. Active Phase Distribution Changes within a Catalyst Particle during Fischer–Tropsch Synthesis as Revealed by Multi-Scale Microscopy. *Catal. Sci. Technol.* **2016**, *6* (12), 4438–4449.

- (105) Reuel, R. C.; Bartholomew, C. H. Effects of Support and Dispersion on the CO Hydrogenation Activity/Selectivity Properties of Cobalt. *J. Catal.* **1984**, *85* (1), 78–88.
- (106) De La Osa, A. R.; De Lucas, A.; Romero, A.; Valverde, J. L.; Sánchez, P. Influence of the Catalytic Support on the Industrial Fischer–Tropsch Synthetic Diesel Production. *Catal. Today* **2011**, *176* (1), 298–302.
- (107) Oh, J. H.; Bae, J. W.; Park, S. J.; Khanna, P. K.; Jun, K. W. Slurry-Phase Fischer–Tropsch Synthesis Using Co/ γ -Al₂O₃, Co/SiO₂ and Co/TiO₂: Effect of Support on Catalyst Aggregation. *Catal. Lett.* **2009**, *130* (3–4), 403–409.
- (108) Eschemann, T. O.; Oenema, J.; de Jong, K. P. Effects of Noble Metal Promotion for Co/TiO₂ Fischer–Tropsch Catalysts. *Catal. Today* **2016**, *261*, 60–66.
- (109) Li, J.; Jacobs, G.; Zhang, Y.; Das, T.; Davis, B. H. Fischer–Tropsch Synthesis: Effect of Small Amounts of Boron, Ruthenium and Rhenium on Co/TiO₂ Catalysts. *Appl. Catal., A* **2002**, *223* (1–2), 195–203.
- (110) Jalama, K.; Kabuba, J.; Xiong, H.; Jewell, L. L. Co/TiO₂ Fischer–Tropsch Catalyst Activation by Synthesis Gas. *Catal. Commun.* **2012**, *17*, 154–159.
- (111) Li, J.; Xu, L.; Keogh, R.; Davis, B. Fischer–Tropsch Synthesis. Effect of CO Pretreatment on a Ruthenium Promoted Co/TiO₂. *Catal. Lett.* **2000**, *70*, 127–130.
- (112) Chai, S. H.; Wang, H. P.; Liang, Y.; Xu, B. Q. Sustainable Production of Acrolein: Gas-Phase Dehydration of Glycerol over Nb₂O₅ catalyst. *J. Catal.* **2007**, *250* (2), 342–349.
- (113) Silva, R. R. C. M.; Schmal, M.; Frety, R.; Dalmon, J. A. Effect of the Support on the Fischer–Tropsch Synthesis with Co/Nb₂O₅ Catalysts. *J. Chem. Soc., Faraday Trans.* **1993**, *89* (21), 3975.
- (114) Mendes, F. M. T.; Uhl, A.; Starr, D. E.; Guimond, S.; Schmal, M.; Kuhlbeck, H.; Shaikhutdinov, S. K.; Freund, H. J. Strong Metal Support Interaction on Co/Niobia Model Catalysts. *Catal. Lett.* **2006**, *111* (1–2), 35–41.
- (115) den Otter, J. H.; Yoshida, H.; Ledesma, C.; Chen, D.; de Jong, K. P. On the Superior Activity and Selectivity of PtCo/Nb₂O₅ Fischer–Tropsch Catalysts. *J. Catal.* **2016**, *340*, 270–275.
- (116) Hernández Mejía, C.; den Otter, J. H.; Weber, J. L.; de Jong, K. P. Crystalline Niobia with Tailored Porosity as Support for Cobalt Catalysts for the Fischer–Tropsch Synthesis. *Appl. Catal., A* **2017**, *548*, 143–149.
- (117) Zhang, Y.; Xiang, H.; Zhong, B.; Wang, Q. ZrO₂-SiO₂ Aerogel Supported Cobalt Catalysts for the Synthesis of Long-Chain Hydrocarbon from Syngas. *Pet. Sci. Technol.* **1999**, *17* (9–10), 981–998.
- (118) Enache, D. I.; Rebours, B.; Roy-Auberger, M.; Revel, R. In Situ XRD Study of the Influence of Thermal Treatment on the Characteristics and the Catalytic Properties of Cobalt-Based Fischer–Tropsch Catalysts. *J. Catal.* **2002**, *205* (2), 346–353.
- (119) Panpranot, J.; Taochaiyaphum, N.; Jongsomjit, B.; Praserthdam, P. Differences in Characteristics and Catalytic Properties of Co Catalysts Supported on Micron- and Nano-Sized Zirconia. *Catal. Commun.* **2006**, *7* (3), 192–197.
- (120) Liu, Y.; Chen, J.; Fang, K.; Wang, Y.; Sun, Y. A Large Pore-Size Mesoporous Zirconia Supported Cobalt Catalyst with Good Performance in Fischer–Tropsch Synthesis. *Catal. Commun.* **2007**, *8* (6), 945–949.
- (121) Soisuwan, S.; Netiworaruksa, B.; Charoendechanukor, C.; Tubcharoen, T.; Panpranot, J.; Praserthdam, P. Characteristics and Catalytic Properties of La-Modified ZrO₂ Supported Cobalt Catalysts in CO Hydrogenation. *React. Kinet., Mech. Catal.* **2011**, *103* (2), 367–378.
- (122) Zhang, X.; Su, H.; Zhang, Y.; Gu, X. Effect of CeO₂ Promotion on the Catalytic Performance of Co/ZrO₂ Catalysts for Fischer–Tropsch Synthesis. *Fuel* **2016**, *184*, 162–168.
- (123) Opanasenko, M. Zeolite Constructor Kit: Design for Catalytic Applications. *Catal. Today* **2018**, *304*, 2–11.
- (124) McMahon, K. C.; Suib, S. L.; Johnson, B. G.; Bartholomew, C. H., Jr. Dispersed Cobalt-Containing Zeolite Fischer–Tropsch Catalysts. *J. Catal.* **1987**, *106* (1), 47–53.
- (125) Bartholomew, C. H. Recent Developments in Fischer–Tropsch Catalysis. *Stud. Surf. Sci. Catal.* **1991**, *64*, 158–224.
- (126) Fornasari, G.; La, T. M. G.; Vaccari, A.; Bednarova, S.; Jiru, P.; Tvaruzkova, Z. Fischer–Tropsch Synthesis on Zeolite Supported Cobalt Catalysts. *Stud. Surf. Sci. Catal.* **1991**, *61* (C), 333–339.
- (127) Udaya, V.; Rao, S.; Gormley, R. J. Bifunctional Catalysis in Syngas Conversions. *Catal. Today* **1990**, *6* (3), 207–234.
- (128) Bessell, S. ZSM-5 as a Support for Cobalt Fischer–Tropsch Catalysts. *Stud. Surf. Sci. Catal.* **1994**, *81* (C), 461–466.
- (129) Ngamcharussrivichai, C.; Liu, X.; Li, X.; Vitidsant, T.; Fujimoto, K. An Active and Selective Production of Gasoline-Range Hydrocarbons over Bifunctional Co-Based Catalysts. *Fuel* **2007**, *86* (1–2), 50–59.
- (130) Dalil, M.; Sohrabi, M.; Royaeae, S. J. Application of Nano-Sized Cobalt on ZSM-5 Zeolite as an Active Catalyst in Fischer–Tropsch Synthesis. *J. Ind. Eng. Chem.* **2012**, *18* (2), 690–696.
- (131) Zola, A. S.; Bidart, A. M. F.; Fraga, A. d. C.; Hori, C. E.; Sousa-Aguiar, E. F.; Arroyo, P. A. Cobalt Supported on Different Zeolites for Fischer–Tropsch Synthesis. *Stud. Surf. Sci. Catal.* **2007**, *167*, 129–134.
- (132) Higgins, J. B.; LaPierre, R. B.; Schlenker, J. L.; Rohrman, A. C.; Wood, J. D.; Kerr, G. T.; Rohrbach, W. J. The Framework Topology of Zeolite Beta. *Zeolites* **1988**, *8* (6), 446–452.
- (133) Pereira, A. L. C.; González-Carballo, J. M.; Pérez-Alonso, F. J.; Rojas, S.; Fierro, J. L. G.; Rangel, M. D. C. Effect of the Mesostucturation of the Beta Zeolite Support on the Properties of Cobalt Catalysts for Fischer–Tropsch Synthesis. *Top. Catal.* **2011**, *54* (1–4), 179–189.
- (134) Espinosa, G.; Domínguez, J. M.; Morales-Pacheco, P.; Tobon, A.; Aguilar, M.; Benítez, J. Catalytic Behavior of Co/(Nano β -Zeolite) Bifunctional Catalysts for Fischer–Tropsch Reactions. *Catal. Today* **2011**, *166* (1), 47–52.
- (135) Nakanishi, M.; Uddin, M. A.; Kato, Y.; Nishina, Y.; Hapipi, A. M. Effects of Preparation Method on the Properties of Cobalt Supported B-Zeolite Catalysts for Fischer–Tropsch Synthesis. *Catal. Today* **2017**, *291*, 124–132.
- (136) Meng, X.; Nawaz, F.; Xiao, F. S. Templating Route for Synthesizing Mesoporous Zeolites with Improved Catalytic Properties. *Nano Today* **2009**, *4* (4), 292–301.
- (137) Xing, C.; Yang, G.; Wu, M.; Yang, R.; Tan, L.; Zhu, P.; Wei, Q.; Li, J.; Mao, J.; Yoneyama, Y.; Tsubaki, N. Hierarchical Zeolite y Supported Cobalt Bifunctional Catalyst for Facilely Tuning the Product Distribution of Fischer–Tropsch Synthesis. *Fuel* **2015**, *148*, 48–57.
- (138) Xing, C.; Yang, G.; Lu, P.; Shen, W.; Gai, X.; Tan, L.; Mao, J.; Wang, T.; Yang, R.; Tsubaki, N. A Hierarchically Spherical Co-Based Zeolite Catalyst with Aggregated Nanorods Structure for Improved Fischer–Tropsch Synthesis Reaction Activity and Isoparaffin Selectivity. *Microporous Mesoporous Mater.* **2016**, *233*, 62–69.
- (139) Lu, P.; Sun, J.; Zhu, P.; Abe, T.; Yang, R.; Taguchi, A.; Vitidsant, T.; Tsubaki, N. Sputtered Nano-Cobalt on H-USY Zeolite for Selectively Converting Syngas to Gasoline. *J. Energy Chem.* **2015**, *24* (5), 637–641.
- (140) Trovarelli, A. Catalytic Properties of Ceria and CeO₂-Containing Materials. *Catal. Rev.: Sci. Eng.* **1996**, *38* (4), 439–520.
- (141) Gnanamani, M. K.; Ribeiro, M. C.; Ma, W.; Shafer, W. D.; Jacobs, G.; Graham, U. M.; Davis, B. H. Fischer–Tropsch Synthesis: Metal-Support Interfacial Contact Governs Oxygenates Selectivity over CeO₂ supported Pt-Co Catalysts. *Appl. Catal., A* **2011**, *393* (1–2), 17–23.
- (142) Yang, Y.; Chiang, K.; Burke, N. Porous Carbon-Supported Catalysts for Energy and Environmental Applications: A Short Review. *Catal. Today* **2011**, *178* (1), 197–205.
- (143) Fu, T.; Li, Z. Review of Recent Development in Co-Based Catalysts Supported on Carbon Materials for Fischer–Tropsch Synthesis. *Chem. Eng. Sci.* **2015**, *135*, 3–20.

- (144) Xiong, H.; Jewell, L. L.; Coville, N. J. Shaped Carbons As Supports for the Catalytic Conversion of Syngas to Clean Fuels. *ACS Catal.* **2015**, *5*, 2640–2658.
- (145) Bezemer, G. L.; van Laak, A.; van Dillen, A. J.; de Jong, K. P. Cobalt Supported on Carbon Nanofibers- a Promising Novel Fischer–Tropsch Catalyst. *Stud. Surf. Sci. Catal.* **2004**, *147*, 259–264.
- (146) Bezemer, G. L.; Bitter, J. H.; Kuipers, H. P. C. E.; Oosterbeek, H.; Holewijn, J. E.; Xu, X.; Kapteijn, F.; van Dillen, A. J.; de Jong, K. P. Cobalt Particle Size Effects in the Fischer–Tropsch Reaction Studied with Carbon Nanofiber Supported Catalysts. *J. Am. Chem. Soc.* **2006**, *128* (12), 3956–3964.
- (147) den Breejen, J. P.; Radstake, P. B.; Bezemer, G. L.; Bitter, J. H.; Frøseth, V.; Holmen, A.; de Jong, K. P. On the Origin of the Cobalt Particle Size Effects in Fischer–Tropsch Catalysis. *J. Am. Chem. Soc.* **2009**, *131* (20), 7197–7203.
- (148) Trépanier, M.; Tavasoli, A.; Dalai, A. K.; Abatzoglou, N. Fischer–Tropsch Synthesis over Carbon Nanotubes Supported Cobalt Catalysts in a Fixed Bed Reactor: Influence of Acid Treatment. *Fuel Process. Technol.* **2009**, *90* (3), 367–374.
- (149) Trépanier, M.; Tavasoli, A.; Dalai, A. K.; Abatzoglou, N. Co, Ru and K Loadings Effects on the Activity and Selectivity of Carbon Nanotubes Supported Cobalt Catalyst in Fischer–Tropsch Synthesis. *Appl. Catal., A* **2009**, *353* (2), 193–202.
- (150) Fu, T.; Lv, J.; Li, Z. Effect of Carbon Porosity and Cobalt Particle Size on the Catalytic Performance of Carbon Supported Cobalt Fischer–Tropsch Catalysts. *Ind. Eng. Chem. Res.* **2014**, *53* (4), 1342–1350.
- (151) Davari, M.; Karimi, S.; Tavasoli, A.; Karimi, A. Enhancement of Activity, Selectivity and Stability of CNTs-Supported Cobalt Catalyst in Fischer–Tropsch via CNTs Functionalization. *Appl. Catal., A* **2014**, *485*, 133–142.
- (152) Xiong, H.; Moyo, M.; Rayner, M. K.; Jewell, L. L.; Billing, D. G.; Coville, N. J. Autoreduction and Catalytic Performance of a Cobalt Fischer–Tropsch Synthesis Catalyst Supported on Nitrogen-Doped Carbon Spheres. *ChemCatChem* **2010**, *2* (5), 514–518.
- (153) Su, D. S.; Perathoner, S.; Centi, G. Nanocarbons for the Development of Advanced Catalysts. *Chem. Rev.* **2013**, *113* (8), 5782–5816.
- (154) Tavasoli, A.; Trépanier, M.; Dalai, A. K.; Abatzoglou, N. Effects of Confinement in Carbon Nanotubes on the Activity, Selectivity, and Lifetime of Fischer–Tropsch Co/Carbon Nanotube Catalysts. *J. Chem. Eng. Data* **2010**, *55* (8), 2757–2763.
- (155) Xing, C.; Yang, G.; Wang, D.; Zeng, C.; Jin, Y.; Yang, R.; Suehiro, Y.; Tsubaki, N. Controllable Encapsulation of Cobalt Clusters inside Carbon Nanotubes as Effective Catalysts for Fischer–Tropsch Synthesis. *Catal. Today* **2013**, *215*, 24–28.
- (156) Karimi, S.; Tavasoli, A.; Mortazavi, Y.; Karimi, A. Cobalt Supported on Graphene - A Promising Novel Fischer–Tropsch Synthesis Catalyst. *Appl. Catal., A* **2015**, *499*, 188–196.
- (157) Eschemann, T. O.; Lamme, W. S.; Manchester, R. L.; Parmentier, T. E.; Cognigni, A.; Rønning, M.; de Jong, K. P. Effect of Support Surface Treatment on the Synthesis, Structure, and Performance of Co/CNT Fischer–Tropsch Catalysts. *J. Catal.* **2015**, *328*, 130–138.
- (158) Ali, S.; Zabidi, N. A. M.; Al-Marri, M. J.; Khader, M. M. Effect of the Support on Physicochemical Properties and Catalytic Performance of Cobalt Based Nano-Catalysts in Fischer–Tropsch Reaction. *Mater. Today Commun.* **2017**, *10*, 67–71.
- (159) Phaahlamohlaka, T. N.; Kumi, D. O.; Dlamini, M. W.; Forbes, R.; Jewell, L. L.; Billing, D. G.; Coville, N. J. Effects of Co and Ru Intimacy in Fischer–Tropsch Catalysts Using Hollow Carbon Sphere Supports: Assessment of the Hydrogen Spillover Processes. *ACS Catal.* **2017**, *7* (3), 1568–1578.
- (160) Aluha, J.; Boahene, P.; Dalai, A.; Hu, Y.; Bere, K.; Braid, N.; Abatzoglou, N. Synthesis and Characterization of Co/C and Fe/C Nanocatalysts for Fischer–Tropsch Synthesis: A Comparative Study Using a Fixed-Bed Reactor. *Ind. Eng. Chem. Res.* **2015**, *54* (43), 10661–10674.
- (161) Aluha, J.; Braid, N.; Dalai, A.; Abatzoglou, N. Low-Temperature Fischer–Tropsch Synthesis Using Plasma-Synthesized Nanometric Co/C and Fe/C Catalysts. *Can. J. Chem. Eng.* **2016**, *94* (8), 1504–1515.
- (162) Aluha, J.; Abatzoglou, N. Synthetic Fuels from 3- φ Fischer–Tropsch Synthesis Using Syngas Feed and Novel Nanometric Catalysts Synthesized by Plasma. *Biomass Bioenergy* **2016**, *95*, 330–339.
- (163) Aluha, J.; Hu, Y.; Abatzoglou, N. Effect of CO Concentration on the α -Value of Plasma-Synthesized Co/C Catalyst in Fischer–Tropsch Synthesis. *Catalysts* **2017**, *7* (3), 69.
- (164) Aluha, J.; Abatzoglou, N. Promotional Effect of Mo and Ni in Plasma-Synthesized Co–Fe/C Bimetallic Nano-Catalysts for Fischer–Tropsch Synthesis. *J. Ind. Eng. Chem.* **2017**, *50*, 199–212.
- (165) Aluha, J.; Abatzoglou, N. Gold-Promoted Plasma-Synthesized Ni-Co-Fe/C Catalyst for Fischer–Tropsch Synthesis. *Gold Bull.* **2017**, *50* (2), 147–162.
- (166) Aluha, J.; Bere, K.; Abatzoglou, N.; Gitzhofer, F. Synthesis of Nano-Catalysts by Induction Suspension Plasma Technology (SPS) for Fischer–Tropsch Reaction. *Plasma Chem. Plasma Process.* **2016**, *36* (5), 1325–1348.
- (167) Zhu, X.; Lu, X.; Liu, X.; Hildebrandt, D.; Glasser, D. Study of Radial Heat Transfer in a Tubular Fischer–Tropsch Synthesis Reactor. *Ind. Eng. Chem. Res.* **2010**, *49* (21), 10682–10688.
- (168) Zhu, X.; Lu, X.; Liu, X.; Hildebrandt, D.; Glasser, D. Heat Transfer Study with and without Fischer–Tropsch Reaction in a Fixed Bed Reactor with TiO₂, SiO₂, and SiC Supported Cobalt Catalysts. *Chem. Eng. J.* **2014**, *247*, 75–84.
- (169) Duong-Viet, C.; Ba, H.; El-Berrichi, Z.; Nhut, J.-M.; Ledoux, M. J.; Liu, Y.; Pham-Huu, C. Silicon Carbide Foam as a Porous Support Platform for Catalytic Applications. *New J. Chem.* **2016**, *40* (5), 4285–4299.
- (170) Nguyen, P.; Pham, C. Innovative Porous SiC-Based Materials: From Nanoscopic Understandings to Tunable Carriers Serving Catalytic Needs. *Appl. Catal., A* **2011**, *391* (1–2), 443–454.
- (171) Liu, Y.; Ersen, O.; Meny, C.; Luck, F.; Pham-Huu, C. Fischer–Tropsch Reaction on a Thermally Conductive and Reusable Silicon Carbide Support. *ChemSusChem* **2014**, *7* (5), 1218–1239.
- (172) Liu, Y.; Edouard, D.; Nguyen, L. D.; Begin, D.; Nguyen, P.; Pham, C.; Pham-Huu, C. High Performance Structured Platelet Millireactor Filled with Supported Cobalt Open Cell SiC Foam Catalyst for the Fischer–Tropsch Synthesis. *Chem. Eng. J.* **2013**, *222*, 265–273.
- (173) de la Osa, A.; Romero, A.; Dorado, F.; Valverde, J.; Sánchez, P. Influence of Cobalt Precursor on Efficient Production of Commercial Fuels over FTS Co/SiC Catalyst. *Catalysts* **2016**, *6* (7), 98.
- (174) Díaz, J. A.; Calvo-Serrano, M.; de la Osa, A. R.; García-Minguillán, A. M.; Romero, A.; Giroir-Fendler, A.; Valverde, J. L. β -Silicon Carbide as a Catalyst Support in the Fischer–Tropsch Synthesis: Influence of the Modification of the Support by a Pore Agent and Acidic Treatment. *Appl. Catal., A* **2014**, *475*, 82–89.
- (175) Liu, Y.; de Tymowski, B.; Vigneron, F.; Florea, I.; Ersen, O.; Meny, C.; Nguyen, P.; Pham, C.; Luck, F.; Pham-Huu, C. Titania-Decorated Silicon Carbide-Containing Cobalt Catalyst for Fischer–Tropsch Synthesis. *ACS Catal.* **2013**, *3* (3), 393–404.
- (176) Moffat, J. B. Phosphates as Catalysts. *Catal. Rev.: Sci. Eng.* **1978**, *18* (2), 199–258.
- (177) Cheung, T. T. P.; Willcox, K. W.; McDaniel, M. P.; Johnson, M. M.; Bronnimann, C.; Frye, J. The Structure of Coprecipitated Aluminophosphate Catalyst Supports. *J. Catal.* **1986**, *102* (1), 10–20.
- (178) Bae, J. W.; Kim, S. M.; Kang, S. H.; Chary, K. V. R.; Lee, Y. J.; Kim, H. J.; Jun, K. W. Effect of Support and Cobalt Precursors on the Activity of Co/AlPO₄ Catalysts in Fischer–Tropsch Synthesis. *J. Mol. Catal. A: Chem.* **2009**, *311* (1–2), 7–16.
- (179) Cavani, F.; Trifirò, F.; Vaccari, A. Hydrotalcite-Type Anionic Clays: Preparation, Properties and Applications. *Catal. Today* **1991**, *11* (2), 173–301.

- (180) Krylova, M. V.; Kulikov, A. B.; Knyazeva, M. I.; Krylova, A. Yu. Cobalt-Containing Catalysts Made from Layered Double Hydroxides for Synthesis of Hydrocarbon from Carbon Monoxide and Hydrogen. *Chem. Technol. Fuels Oils* **2008**, *44* (5), 339–343.
- (181) Tsai, Y.-T.; Mo, X.; Campos, A.; Goodwin, J. G.; Spivey, J. J. Hydrotalcite Supported Co Catalysts for CO Hydrogenation. *Appl. Catal., A* **2011**, *396* (1–2), 91–100.
- (182) Di Fronzo, A.; Pirola, C.; Comazzi, A.; Galli, F.; Bianchi, C. L.; Di Michele, A.; Vivani, R.; Nocchetti, M.; Bastianini, M.; Boffito, D. Co-Based Hydrotalcites as New Catalysts for the Fischer–Tropsch Synthesis Process. *Fuel* **2014**, *119*, 62–69.
- (183) Chughtai, A. H.; Ahmad, N.; Younus, H. A.; Laypkov, A.; Verpoort, F. Metal–organic Frameworks: Versatile Heterogeneous Catalysts for Efficient Catalytic Organic Transformations. *Chem. Soc. Rev.* **2015**, *44* (19), 6804–6849.
- (184) Liu, J.; Chen, L.; Cui, H.; Zhang, J.; Zhang, L.; Su, C.-Y. Applications of Metal–organic Frameworks in Heterogeneous Supramolecular Catalysis. *Chem. Soc. Rev.* **2014**, *43* (16), 6011–6061.
- (185) Isaeva, V. I.; Eliseev, O. L.; Kazantsev, R. V.; Chernyshev, V. V.; Davydov, P. E.; Saifutdinov, B. R.; Lapidus, A. L.; Kustov, L. M. Fischer–Tropsch Synthesis over MOF-Supported Cobalt Catalysts (Co@MIL-53(Al)). *Dalton Trans.* **2016**, *45* (30), 12006–12014.
- (186) Qiu, B.; Yang, C.; Guo, W.; Xu, Y.; Liang, Z.; Ma, D.; Zou, R. Highly Dispersed Co-Based Fischer–Tropsch Synthesis Catalysts from Metal–organic Frameworks. *J. Mater. Chem. A* **2017**, *5* (17), 8081–8086.
- (187) Sadeqzadeh, M.; Karaca, H.; Safonova, O. V.; Fongarland, P.; Chambrey, S.; Roussel, P.; Griboval-Constant, A.; Lacroix, M.; Curulla-Ferré, D.; Luck, F.; Khodakov, A. Y. Identification of the Active Species in the Working Alumina-Supported Cobalt Catalyst under Various Conditions of Fischer–Tropsch Synthesis. *Catal. Today* **2011**, *164* (1), 62–67.
- (188) Pei, Y.; Li, Z.; Li, Y. Highly Active and Selective Co-Based Fischer–Tropsch Catalysts Derived from Metal–organic Frameworks. *AIChE J.* **2017**, *63* (7), 2935–2944.
- (189) Banhart, J. Manufacture, Characterisation and Application of Cellular Metals and Metal Foams. *Prog. Mater. Sci.* **2001**, *46* (6), 559–632.
- (190) Park, J. C.; Roh, N. S.; Chun, D. H.; Jung, H.; Yang, J. I. Cobalt Catalyst Coated Metallic Foam and Heat-Exchanger Type Reactor for Fischer–Tropsch Synthesis. *Fuel Process. Technol.* **2014**, *119*, 60–66.
- (191) Harmel, J. Synthèse de Nano-Catalyseurs Hybrides à Base de Cobalt Pour La Catalyse Fischer–Tropsch. Ph.D. Thesis, University of Toulouse, Toulouse, France, 2016.
- (192) Narasaraju, T. S. B.; Phebe, D. E. Some Physico-Chemical Aspects of Hydroxylapatite. *J. Mater. Sci.* **1996**, *31* (1), 1–21.
- (193) Pham Minh, D.; Galera Martínez, M.; Nzihou, A.; Sharrock, P. Thermal Behavior of Apatitic Calcium Phosphates Synthesized from Calcium Carbonate and Orthophosphoric Acid or Potassium Dihydrogen Orthophosphate. *J. Therm. Anal. Calorim.* **2013**, *112* (3), 1145–1155.
- (194) Joris, S. J.; Amberg, C. H. Nature of Deficiency in Nonstoichiometric Hydroxyapatites. II. Spectroscopic Studies of Calcium and Strontium Hydroxyapatites. *J. Phys. Chem.* **1971**, *75* (20), 3172–3178.
- (195) Tsuchida, T.; Kubo, J.; Yoshioka, T.; Sakuma, S.; Takeguchi, T.; Ueda, W. Reaction of Ethanol over Hydroxyapatite Affected by Ca/P Ratio of Catalyst. *J. Catal.* **2008**, *259* (2), 183–189.
- (196) Wakamura, M.; Kandori, K.; Ishikawa, T. Surface Structure and Composition of Calcium Hydroxyapatites Substituted with Al(III), La(III) and Fe(III) Ions. *Colloids Surf., A* **2000**, *164* (2–3), 297–305.
- (197) Fihri, A.; Len, C.; Varma, R. S.; Solhy, A. Hydroxyapatite: A Review of Syntheses, Structure and Applications in Heterogeneous Catalysis. *Coord. Chem. Rev.* **2017**, *347*, 48–76.
- (198) Boukha, Z.; Kacimi, M.; Ziyad, M.; Ensuque, A.; Bozon-Verduraz, F. Comparative Study of Catalytic Activity of Pd Loaded Hydroxyapatite and Fluoroapatite in Butan-2-ol Conversion and Methane Oxidation. *J. Mol. Catal. A: Chem.* **2007**, *270* (1–2), 205–213.
- (199) Jun, J. H.; Lee, T.-J.; Lim, T. H.; Nam, S.-W.; Hong, S.-A.; Yoon, K. J. Nickel–Calcium Phosphate/Hydroxyapatite Catalysts for Partial Oxidation of Methane to Syngas: Characterization and Activation. *J. Catal.* **2004**, *221* (1), 178–190.
- (200) Rêgo De Vasconcelos, B.; Zhao, L.; Sharrock, P.; Nzihou, A.; Pham Minh, D. Catalytic Transformation of Carbon Dioxide and Methane into Syngas over Ruthenium and Platinum Supported Hydroxyapatites. *Appl. Surf. Sci.* **2016**, *390*, 141–156.
- (201) Rêgo de Vasconcelos, B.; Pham Minh, D.; Sharrock, P.; Nzihou, A. Regeneration Study of Ni/Hydroxyapatite Spent Catalyst from Dry Reforming. *Catal. Today* **2018**, *310*, 107–115.
- (202) Boukha, Z.; Kacimi, M.; Pereira, M. F. R.; Faria, J. L.; Figueiredo, J. L.; Ziyad, M. Methane Dry Reforming on Ni Loaded Hydroxyapatite and Fluoroapatite. *Appl. Catal., A* **2007**, *317* (2), 299–309.
- (203) Phan, T. S.; Sane, A. R.; Rêgo de Vasconcelos, B.; Nzihou, A.; Sharrock, P.; Grouset, D.; Pham Minh, D. Hydroxyapatite Supported Bimetallic Cobalt and Nickel Catalysts for Syngas Production from Dry Reforming of Methane. *Appl. Catal., B* **2018**, *224*, 310–321.
- (204) Venugopal, A.; Scurrall, M. S. Hydroxyapatite as a Novel Support for Gold and Ruthenium Catalysts: Behaviour in the Water Gas Shift Reaction. *Appl. Catal., A* **2003**, *245* (1), 137–147.
- (205) Yaakob, Z.; Hakim, L.; Kumar, M.; Ismail, M.; Daud, W. Short Communication: Hydroxyapatite Supported Nickel Catalyst for Hydrogen Production. *Am. J. Sci. Ind. Res.* **2010**, *1* (2), 122–126.
- (206) Mori, K.; Hara, T.; Mizugaki, T.; Ebitani, K.; Kaneda, K. Hydroxyapatite-Supported Palladium Nanoclusters: A Highly Active Heterogeneous Catalyst for Selective Oxidation of Alcohols by Use of Molecular Oxygen. *J. Am. Chem. Soc.* **2004**, *126* (34), 10657–10666.